CBSE EXAMINATION PAPER-2008

NEET-ZOOLOGY

(Solved)

Time allowed : 3 hours | Maximum Marks : 799

General Instructions :

Read the following instructions carefully and follow them :

i. This question paper contains **200 questions.** All questions are **compulsory.**

ii. This question paper is divided into **1 sections.**

iii. **Section A** – questions number **1 to 200** are multiple choice questions

iv. There is no overall choice given in the question paper. However, an internal choice has been provided in few questions.

v. Use of calculator is NOT allowed.

Section A

Question 1.

Which two of the following five physical parameters have the same dimensions?

(A) Energy density

(B) Refractive index

(C) Dielectric constant

(D) Young's modulus

(E) Magnetic field

[4 Marks]

(A) (A) and (E)

(B) (A) and (D)

(C) (C) and (E)

(D) (B) and (D)

Explanation:

Energy density = Energy / Volume

⇒ $[ML^2 T^{-2}] / [L^3] = [ML^{-1} T^{-2}]$

Refractive index and dielectric constant have no dimensions.

Young's Modulus = $F / A \times l / \Delta l$

⇒ $[MLT^{-2}] / [L^2] . [L] / [L] = [ML^{-1}T^{-2}]$·

Magnetic field, $B = F / i\ell = [MLT^{-2}] / [A][L]$

$= [MT^{-2}A^{-1}]$

Question 2.

If the error in the measurement of radius of a sphere is 2%, then the error in the determination of volume of the sphere will be:

[4 Marks]

(A) 6%

(B) 2%

(C) 4%

(D) 8%

Explanation:

Given, error in the measurement of radius of a sphere $\Delta r / r \times 100 = 2\%$

Volume of the sphere $V = (4 / 3) \pi r^3$

∴ Error in the volume $\Delta V / V \times 100$

$= \pm 3 . \Delta r / r \times 100 = 3 \times 2\% = \pm 6\%$

Question 3.

The distance travelled by a particle starting from rest and moving with an acceleration $(4 / 3)\ ms^{-2}$, in the third second is:

[4 Marks]

(A) 6m

(B) 4m

(C) (19 / 3) m

(D) (10 / 3)m

Explanation:

Distance travelled in the nth second is given by $S_n = u + (a / 2)(2n - 1)$

Given $u = 0$, $a = (4 / 3)\ ms^{-2}$, $n = 3$

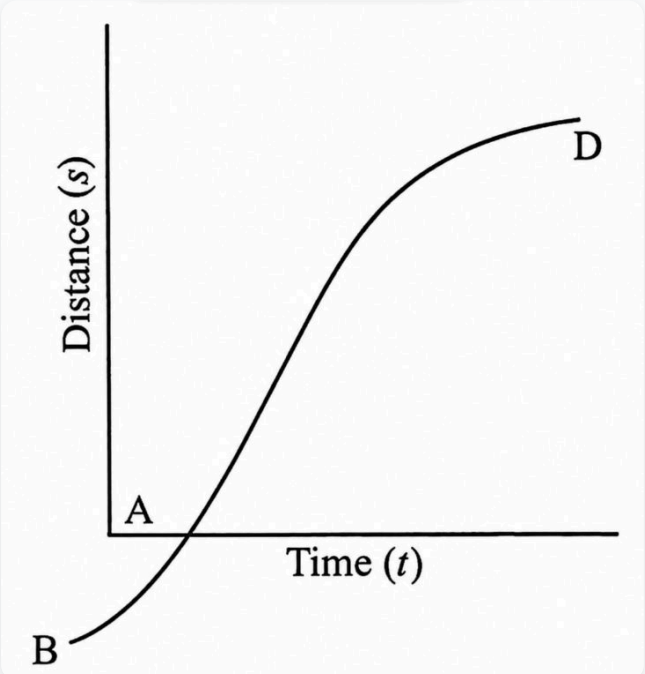
$\therefore S_n = 0 + ((4 / 3 \times 2)(2 \times 3 - 1)) = 4 / 6 \times 5 = (10 / 3)$ m

Question 4.

A particle shows distance–time curve as given in this figure.

The maximum instantaneous velocity of the particle is around the point:

[4 Marks]



(A) A

(B) D

(C) B

(D) C

Explanation:

The slope of the graph (ds / dt) is maximum at C and hence the instantaneous velocity is maximum at C.

Instantaneous velocity,

$$\vec{v}_{ins} = \lim_{\Delta t \to 0} \frac{\Delta \vec{s}}{\Delta t} = \frac{d\vec{s}}{dt}$$

Question 5.

A particle moves in a straight line with a constant acceleration. It changes its velocity from 10 ms^{-1} to 20 ms^{-1} while passing through a distance 135 m in t second. The value of t is:

[4 Marks]

(A) 9

(B) 10

(C) 12

(D) 1.8

Explanation:

Initial velocity, u = 10 ms^{-1}

Final velocity, v = 20 ms^{-1}

Distance, s = 135 m

Let, acceleration = a

Using the formula, $v^2 = u^2 + 2as$

$a = v^2 - u^2 / 2s = ((20)^2 - (10)^2 / 2 \times 135) = 400 - 100 / 2 \times 135$

or, $a = 300 / 2 \times 135 = 150 / 135 = (10 / 9)\ ms^{-2}$

Now, using the relation, $v = u + at$

$t = v - u / a = (20 - 10 / (10 / 9)) = 10 / 10 \times 9\ sec = 9\ sec.$

Question 6.

A particle of mass m is projected with velocity v making an angle of 45° with the horizontal. When the particle lands on the level ground the magnitude of the change in its momentum will be:

[4 Marks]

(A) 2 mv

(B) mv / √2

(C) zero

(D) mv√2

Explanation:

The momentum along a-axis remains unchanged

Clearly, change in momentum along x-axis = $mv\cos\theta - mv\cos\theta = 0$

Momentum changed only in vertical direction or y-axis.

So, $\Delta P = P_{vertical}$

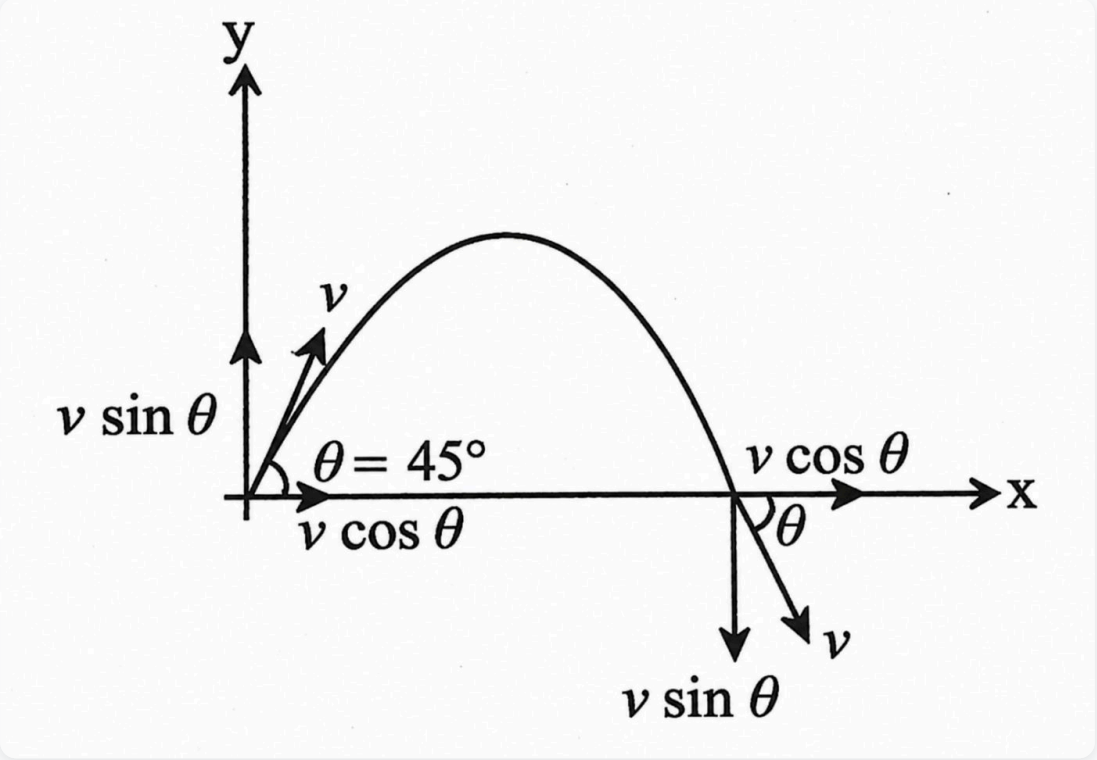
$\Rightarrow P_{final} = P_{initial}$

$= mv\sin\theta - (-mv\sin\theta) = 2mv\sin\theta$

$= 2mv \times \sin 45°$

$= 2mv \times 1 / \sqrt{2} = \sqrt{2}mv$

Hence, resultant change in momentum = √2mv



Question 7.

Sand is being dropped on a conveyor belt at the rate of M kg/s. The force necessary to keep the belt moving with a constant velocity of v m/s will be:

[4 Marks]

(A) zero

(B) 2 Mv newton

(C) (Mv / 2) newton

(D) Mv newton

Explanation:

F= d(Mv) / dt = M(dv / dt) + v (dM / dt)

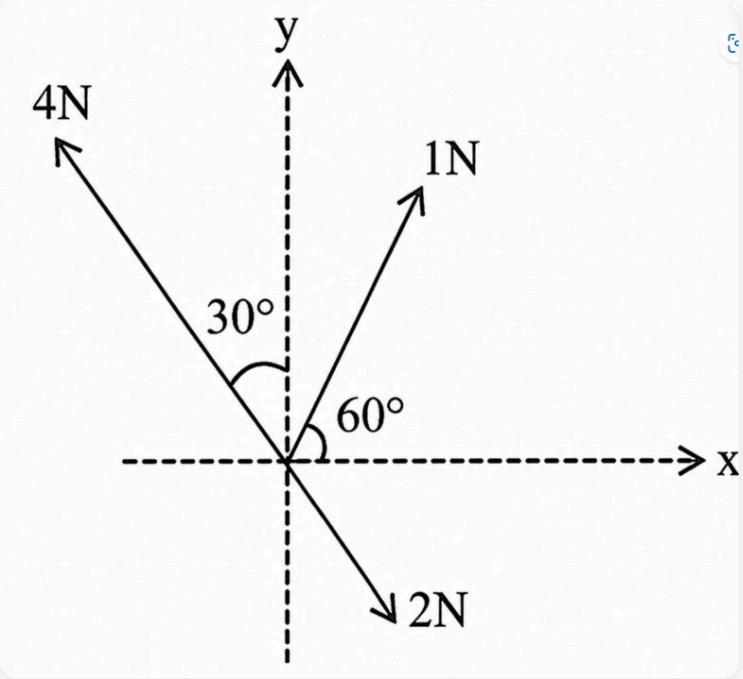
∴ v is constant,

∴ F = v (dM / dt) But (dM / dt) = Mkg/s ∴F=vM newton.

Question 8.

Three forces acting on a body are shown in the figure. To have the resultant force only along the y-direction, the magnitude of the minimum additional force needed is:

[4 Marks]



(A) √3 N

(B) 0.5 N

(C) (√3 / 4) N

(D) 1.5 N

Explanation:

The Components of 1 N and 2 N forces along + x-axis

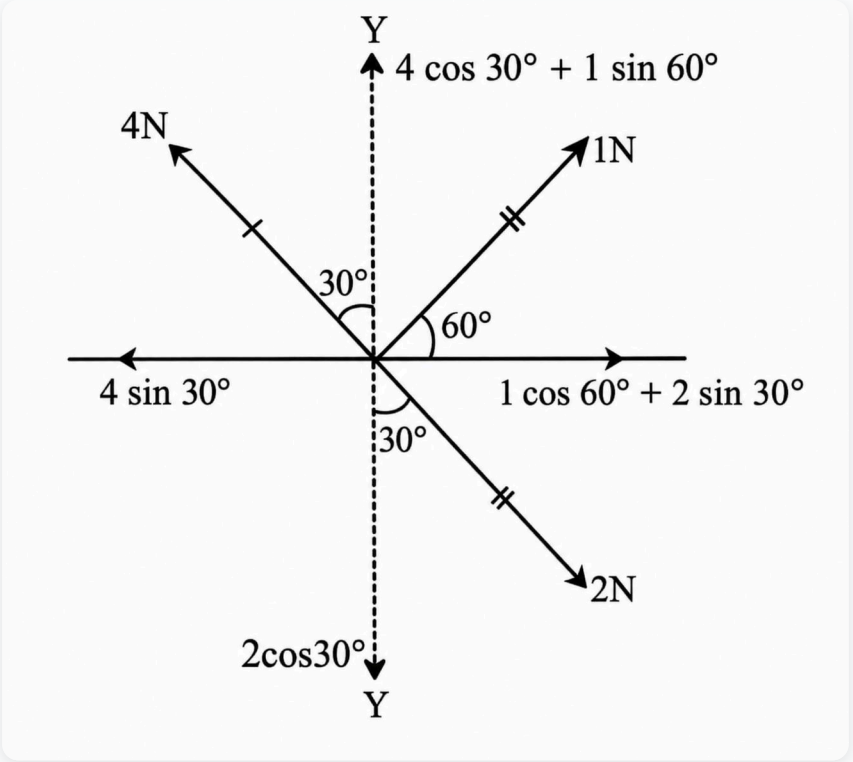
=1 cos 60° + 2 sin 30°

=1× 1 / 2 + 2× 1 / 2 = 1 / 2 + 1 = 3 / 2 = 1.5N

The component of 4 N force along –x-axis = 4 sin 30°

= 4 × 1 / 2 = 2 N.

Therefore, if a force of 0.5 N is applied along +x-axis, the resultant force along x-axis will become zero and the resultant force will be obtained only along y-axis.



Question 9.

Water falls from a height of 60 m at the rate of 15 kg/s to operate a turbine. The losses due to frictional force are 10% of energy. How much power is generated by the turbine? (g = 10 m/s²)

[4 Marks]

(A) 10.2 kW

(B) 7.0 kW

(C) 8.1 kW

(D) 12.3 kW

Explanation:

Given, h = 60 m, g = 10 ms^{-2},

Rate of flow of water = 15 kg/s

∴ Power of the falling water

= 15 kgs^{-1} × 10 ms^{-2} × 60 m = 9000 watt.

Loss in energy due to friction

= 9000 × 10 / 100 = 900 watt.

∴ Power generated by the turbine

= (9000 – 900) watt = 8100 watt = 8.1 kW

Question 10.

A shell of mass 200 gm is ejected from a gun of mass 4 kg by an explosion that generates 1.05 kJ of energy. The initial velocity of the shell is:

[4 Marks]

(A) 120 ms^{-1}

(B) 100 ms^{-1}

(C) 40 ms^{-1}

(D) 80 ms^{-1}

Explanation:

Let the initial velocity of the shell be v, then by the conservation of momentum $mv = Mv'$

where v' = velocity of gun.

$\therefore v' = (m / M) v$

Now, total K.E. = $(1 / 2)mv^2 + (1 / 2)Mv'^2$

$= (1 / 2)mv^2 + ((1 / 2) M) (m / M)^2 v^2 = (1 / 2)mv^2 [1 + m / M]$

$= (1 / 2 \times 0.2) (1 + 0.2 / 4) v^2 = (0.1 \times 1.05) v^2$

But total K.E. $=1.05$ kJ $=1.05 \times 10^3$ J

$\therefore 1.05 \times 10^3 = 0.1 \times 1.05 \times v^2$

$\Rightarrow v^2 = 1.05 \times 10^3 / 0.1 \times 1.05 = 10^4$

$\therefore v = 10^2 = 100\ ms^{-1}$

Question 11.

The ratio of radii of gyration of a circular ring and a circular disc, of the same mass and radius, about an axis passing through their centres and perpendicular to their planes are:

[4 Marks]

(A) 3 : 2

(B) √2 : 1

(C) 2 : 1

(D) 1 : √2

Explanation:

∵ $I = MK^2$ ∴ $K = \sqrt{(I / M)}$

$I_{ring} = MR^2$ and $I_{disc} = (1 / 2) MR^2$

$K_1 / K_2 = \sqrt{(I_1 / I_2)} = \sqrt{(MR^2 / (MR^2/2))} = \sqrt{(2:1)}$

Question 12.

A thin rod of length L and mass M is bent at its midpoint into two halves so that the angle between them is 90°. The moment of inertia of the bent rod about an axis passing through the bending point and perpendicular to the plane defined by the two halves of the rod is:

[4 Marks]

(A) $ML^2 / 12$

(B) $ML^2 / 24$

(C) $ML^2 / 6$

(D) $\sqrt{2}ML^2 / 24$

Explanation:

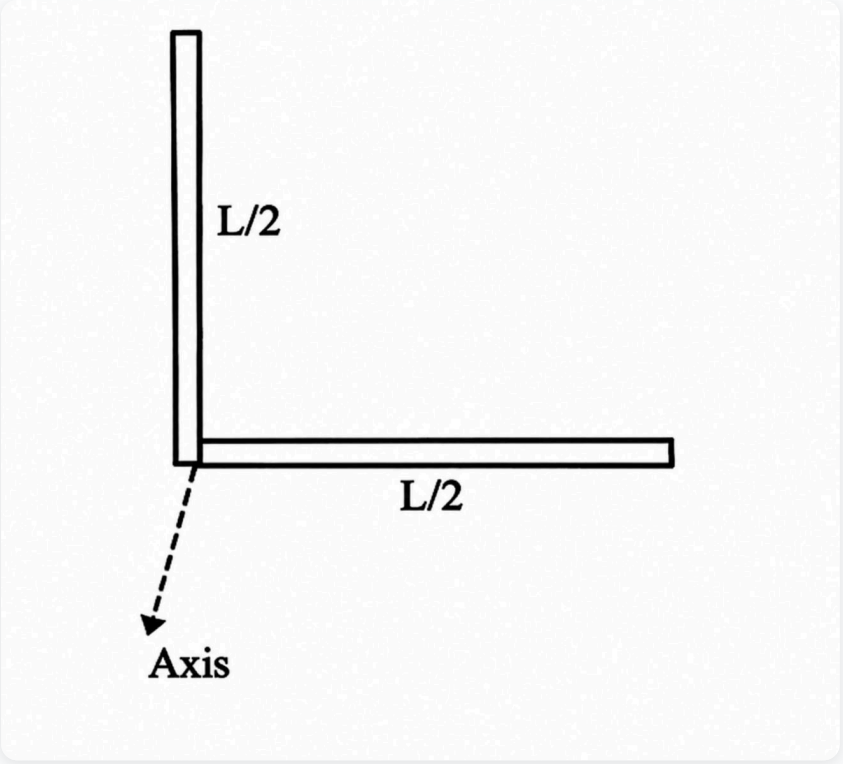
Mass of each part = M / 2

Length of each part = L / 2

Total M.I. = Sum of M.I.s of both parts

$= (M / 2) (L / 2)^2 \times 1 / 3 + (M / 2) (L / 2)^2 \times 1 / 3$

$=2\times M / 2 \times L^2 / 4 \times 1 / 3 = ML^2 / 12$



Question 13.

A roller coaster is designed such that riders experience “weightlessness” as they go round the top of a hill whose radius of curvature is 20 m. The speed of the car at the top of the hill is between:

[4 Marks]

(A) 14 m/s and 15 m/s

(B) 16 m/s and 17 m/s

(C) 15 m/s and 16 m/s

(D) 13 m/s and 14 m/s

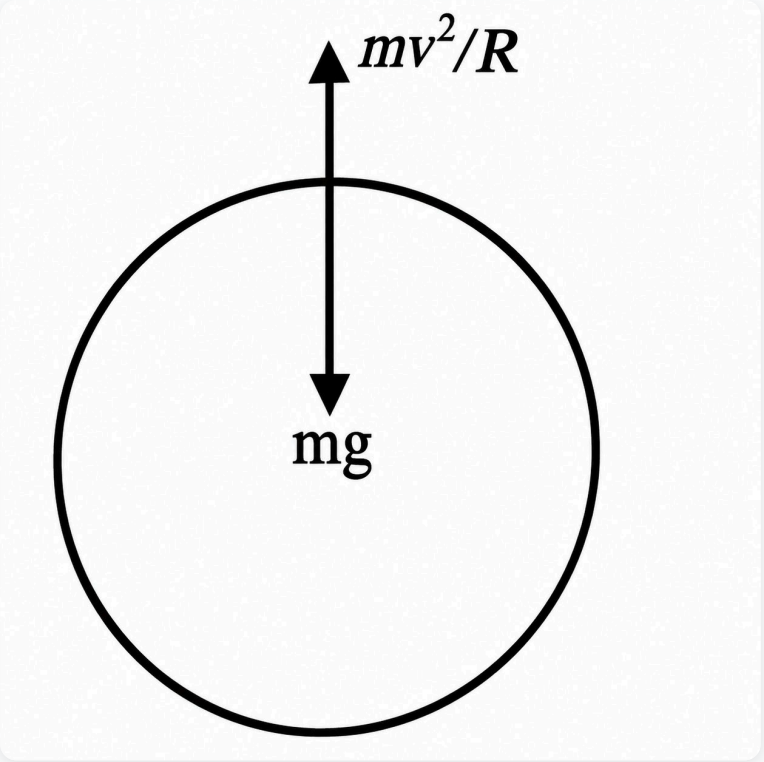
Explanation:

For the riders to experience weightlessness at the top of the hill, the weight of the rider must be balanced by the centripetal force.

i.e., $mg = mv^2 / R$

$\Rightarrow v = \sqrt{gR} = \sqrt{10 \times 20} = 14.1\ ms^{-1}$

Hence, the speed of the car should be between 14 ms^{-1} an 15 ms^{-1}



Question 14.

On a new scale of temperature (which is linear) and called the W scale, the freezing and boiling points of water are 39° W and 239° W respectively. What will be the temperature on the new scale, corresponding to a temperature of 39° C on the Celsius scale?

[4 Marks]

(A) 200° W

(B) 139° W

(C) 117° W

(D) 78° W

Explanation:

For different temperature scales, we have

x − L.F.P / U.F.P. − L.F.P. =constant

Where L.F.P ⇒ Lower Fixed Point

U.F.P ⇒ Upper Fixed Point

Where x is the measurement at that scale. Here, if C and W be the measurements on Celsius and W scale then,

C - 0 / 100 - 0 = W - 39 / 239 - 39 (∴ C = 39°C)

⇒ W = (39 × 200 / 100) + 39 = 78 + 39 = 117° W

Question 15.

An electric kettle takes 4 A current at 220 V. How much time will it take to boil 1 kg of water from temperature 20° C? The temperature of boiling water is 100° C.

[4 Marks]

(A) 12.6 min

(B) 4.2 min

(C) 6.3 min

(D) 8.4 min

Explanation:

Heat required to raise the temperature of 1 kg water from 20°C to 100°C is given by

$Q = ms\Delta\theta = (1 \times 4200 \times 80)$ J

Power of kettle (P) = VI = (220 × 4) W

Time taken=Q / P = ((1 × 4200 × 80) / (220 × 4))

= 381.81 sec =6.36 min

Question 16.

If Q, E and W denote respectively the heat added, change in internal energy and the work done in a closed cyclic process, then:

[4 Marks]

(A) W = 0

(B) Q = W = 0

(C) E = 0

(D) Q = 0

Explanation:

In a cyclic process, the initial state coincides with the final state. Hence, the change in internal energy is zero, as it depends only on the initial and final states.

Question 17.

At 10°C the value of the density of a fixed mass of an ideal gas divided by its pressure is x. At 110°C this ratio is:

[4 Marks]

(A) x

(B) (383 / 283)x

(C) (283 / 383)x

(D) (10 / 110)x

Explanation:

Let the mass of the gas be m.

At a fixed temperature and pressure, volume is fixed.

Density of the gas, $\rho = m / V$

Now $\rho / P = m / nRT \Rightarrow m / nRT = x$ (By question)

$\Rightarrow xT = \text{constant} \Rightarrow x_1T_1 = x_2T_2$

$\Rightarrow x_2 \Rightarrow x_1T_1 / T_2 = (283 / 383)x$

[given $T_1 = 283K$]

[$T_2 = 383K$]

Question 18.

Two simple harmonic motions of angular frequency 100 and 1000 rad s^{-1} have the same displacement amplitude. The ratio of their maximum acceleration is:

[4 Marks]

(A) $1 : 10^3$

(B) 1 : 10

(C) $1 : 10^4$

(D) $1 : 10^2$

Explanation:

Maximum acceleration of a particle in the simple harmonic motion is directly proportional to the square of angular frequency

i.e., $a \propto \omega^2$

$\therefore a_1 / a_2 = \omega_1^2 / \omega_2^2 = (100)^2 / (1000)^2 = 1 / 100$

$\Rightarrow a_1 : a_2 = 1 : 10^2$.

Question 19.

A point performs simple harmonic oscillation of period T and the equation of motion is given by $x = a \sin(\omega t + \pi / 6)$. After the elapse of what fraction of the time period the velocity of the point will be equal to half of its maximum velocity?

[4 Marks]

(A) T / 8

(B) T / 3

(C) T / 6

(D) T / 12

Explanation:

We have $x = a \sin (\omega t + \pi / 6)$

$\therefore$ Velocity, $v = dx / dt = a\omega \cos (\omega t + \pi / 6)$

Maximum velocity = $a\omega$

According to question,

$a\omega / 2 = a\omega \cos (\omega t + \pi / 6)$

or, $\cos (\omega t + \pi / 6) = 1 / 2 = \cos 60° \text{ or } \cos \pi / 3$

$\Rightarrow \omega t + \pi / 6 = \pi / 3$

$\omega t = \pi / 3 - \pi / 6$ or, $\omega t = \pi / 6$

or, $(2\pi / T) \cdot t = \pi / 6 \Rightarrow t = T / 12$

Question 20.

The wave described by $y=0.25\sin(10\pi x-2\pi t)$,

where x and y are in meters and t in seconds, is a wave travelling along the:

[4 Marks]

(A) −ve x direction with frequency 1 Hz.

(B) +ve x direction with frequency π Hz and wavelength λ = 0.2 m.

(C) +ve x direction with frequency 1 Hz and wavelength λ = 0.2 m.

(D) −ve x direction with amplitude 0.25 m and wavelength λ = 0.2 m

Explanation:

$y=0.25\sin(10\pi x-2\pi t)$

Comparing this equation with the standard wave equation

$y=a\sin(kx-\omega t)$

We get, $k=10\pi$

$\Rightarrow 2\pi / \lambda = 10\pi \Rightarrow \lambda = 0.2$ m

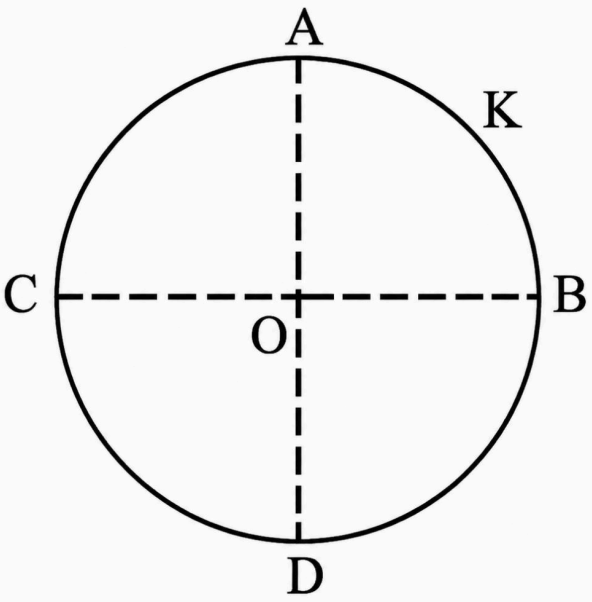
And $\omega = 2\pi$ or, $2\pi\nu = 2\pi \Rightarrow \nu = 1$ Hz

The sign inside the bracket is negative, hence the wave travels in +ve x-direction.

Question 21.

A thin conducting ring of radius R is given a charge +Q. The electric field at the centre O of the ring due to the charge on the part AKB of the ring is E. The electric field at the centre due to the charge on the part ACDB of the ring is

[4 Marks]



(A) E along KO

(B) E along KO

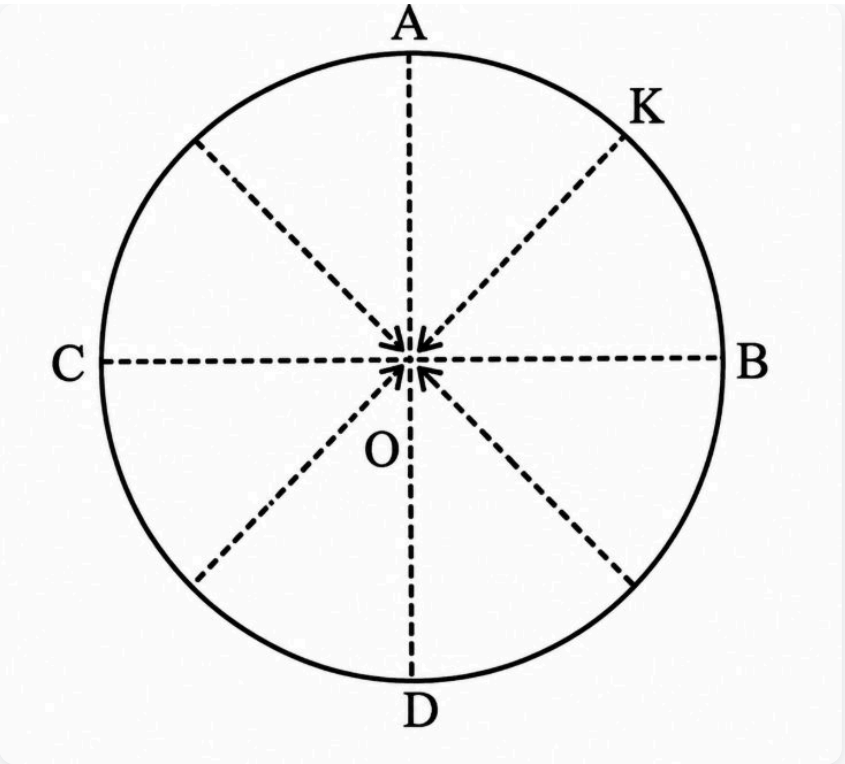
(C) E along OK

(D) 3 E along OK

Explanation:

By the symmetry of the figure, the electric fields at O due to the portions AC and BD are equal in magnitude and opposite in direction. So, they cancel each other.

Similarly, the field at O due to CD and AKB are equal in magnitude but opposite in direction. Therefore, the electric field at the centre due to the charge on the part ACDB is E along OK.



Question 22.

The electric potential at a point in free space due to a charge Q coulomb is $Q \times 10^{11}$ volts. The electric field at that point is

[4 Marks]

(A) $12\pi\varepsilon_0\ Q \times 10^{22}$ volt/m

(B) $4\pi\varepsilon_0\ Q \times 10^{22}$ volt/m

(C) $12\pi\varepsilon_0\ Q \times 10^{20}$ volt/m

(D) $4\pi\varepsilon_0\ Q \times 10^{20}$ volt/m

Explanation:

Given that, $V = Q \times 10^{11}$ volts

Electric potential at point is given by

$V = 1 / 4\pi\ \varepsilon_0\ .\ Q / r$ or, $Q \times 10^{11} = 1 / 4\pi\ \varepsilon_0\ .\ Q / r$

$\Rightarrow r = 1 / 4\pi\ \varepsilon_0 \times 10^{-11}$ m.

As $|E| = |V| / r = (Q \times 10^{11} / (1 / 4\pi\ \varepsilon_0\ ).10^{-11})$

$= 4\pi\ \varepsilon_0\ Q \times 10^{22}$ volt m^{-1}.

Question 23.

A parallel plate capacitor has a uniform electric field E in the space between the plates. If the distance between the plates is d and area of each plate is A, the energy stored in the capacitor is:

[4 Marks]

(A) $(1 / 2) \varepsilon_0 E^2$

(B) E^2Ad / ε_0

(C) $\varepsilon_0 EAd$

(D) $(1 / 2) \varepsilon_0 E^2 Ad$

Explanation:

The energy stored by a capacitor

$U = (1 / 2) CV^2$ .....(i)

V is the p.d. between two plates of the capacitor

Potential difference V = E.d.

The capacitance of the parallel plate capacitor

$C = A\varepsilon_0 / d$

Substituting the value of C in equation (i):

$U = (1 / 2) (A\varepsilon_0 / d)(Ed)^2 = (1 / 2) A\varepsilon_0E^2d$

Question 24.

A wire of a certain material is stretched slowly by ten per cent. Its new resistance and specific resistance become respectively:

[4 Marks]

(A) 1.21 times, same

(B) 1.1 times, 1.1 times

(C) both remain the same

(D) 1.2 times, 1.3 times

Explanation:

Resistance of a wire is given by R= ρ (l / a)

If the length is increased by 10 % then new

length l' = l + (1 / 10) = (11 / 10)l

In that case, area of cross section of wire would decrease by 10 %

∴ New area of cross section

A′= A - A / 10 = (9 / 10) A

∴ R′ = ρ (ℓ′ / A′) = ρ ((11 / 10)l / (9 / 10)) A

R′ = (1 1 / 9) ρ (l / R) R′=1.21R

The new resistance increases by 1.21 times. The specific resistance (resistivity) remains unchanged as it depends on the nature of the material of the wire.

Question 25.

Bacterial leaf blight of rice is caused by a species of

[4 Marks]

(A) Erwinia

(B) Alternaria

(C) Pseudomonas

(D) Xanthomonas

Explanation:

Bacterial leaf blight of rice is caused by a species *Xanthomonas*. Mature rice plants are infected by these bacteria, lesion begin as water-soaked stripes on the leaf blades and eventually increase in length and width, becoming yellow to grayish-white until the entire leaf dries up.

Question 26.

In the light of recent classification of living organisms into three domains of life (bacteria, archaea and eukarya), which one of the following statements is true about archaea?

[4 Marks]

(A) Archaea completely differ from prokaryotes

(B) Archaea have some novel features that are absent in other prokaryotes and eukaryotes

(C) Archaea completely differ form both prokaryotes and eukaryotes

(D) Archaea resemble eukarya in all respects

Explanation:

A domain of prokaryotic organisms containing the archaeobacteria including the methanogens, which produce methane; the thermoacidophilic bacteria, which live in extremely hot and acidic environments & and the halophilic bacteria, which can only function at high salt concentrations, are abundant in the world's oceans.

Question 27.

Thermococcus, Methanococcus and Methan-obacterium exemplify:

[4 Marks]

(A) Bacteria whose DNA is relaxed or positively supercoiled but which have a cytoskeleton as well as mitochondria

(B) Bacteria that contain a cytoskeleton and ribosomes

(C) Archaebacteria that contain protein homologous to eukaryotic core histones

(D) Archaebacteria that lack any histones resembling those found in eukaryotes but whose DNA is negatively supercoiled

Explanation:

Thermococcus, Methanococcus and Methanobacterium exemplify archaebacteria that contain protein homologous to eukaryotic core histones.

Question 28.

In which one of the following, male and female gametophytes do not have free living independent existence?

[4 Marks]

(A) Polytrichum

(B) Cedrus

(C) Funaria

(D) Pteris

Explanation:

Male and female gametophytes of Cedrus do not have free-living independent existence. Cedrus belongs to conifer.

Question 29.

Which one of the following is heterosporous?

[4 Marks]

(A) Salvinia

(B) Dryopteris

(C) Equisetum

(D) Adiantum

Explanation:

Salvinia is an heterosporous aquatic fern with both annual and perennial species. The heterospores have two types of spores — microspores and megaspores.

Question 30.

Select one of the following pairs of important features distinguishing Gnetum from Cycas and Pinus and showing affinities with angiosperms

[4 Marks]

(A) absence of resin duct and leaf venation

(B) perianth and two integuments

(C) embryo development and apical meristem

(D) presence of vessel elements and absence of archegonia

Explanation:

The important features distinguishing Gnetum from Cycas and Pinus and showing affinities with angiosperms are the presence of vessel elements and absence of archegonia.

Question 31.

Ascaris is characterized by

[4 Marks]

(A) presence of true coelom and metamerism (metamerisation)

(B) presence of true coelom but absence of metamerism

(C) presence of neither true coelom nor metamerism

(D) absence of true coelom but presence of metamerism

Explanation:

Ascaris is a common parasite found in the large intestine of man. It is worldwide in distribution. The number of worms may be 500 or more in a single host. Ascaris is characterized by the presence of neither true coelom nor metamerism.

Question 32.

Which one of the following is NOT a characteristic of phylum Annelida?

[4 Marks]

(A) Ventral nerve cord

(B) Pseudocoelom

(C) Segmentation

(D) Closed circulatory system

Explanation:

Characteristic of phylum Annelida is pseudocoelomate. Pseudocoelomate is any invertebrate animal whose body cavity is a pseudocoel, a cavity between the gut and the outer body wall derived from a persistent blastocoel, rather than a true coelom. Pseudocoelomate animals include the Rotifera and Nematoda.

Question 33.

Which one of the following phyla is correctly matched with its two general characteristics?

[4 Marks]

(A) Mollusca - Normally oviparous and development through a trochophore or veliger larva

(B) Echinodermata - Pentamerous radial symmetry and mostly internal fertilization

(C) Arthropoda – Body divided into head, thorax and abdomen and respiration by tracheae

(D) Chordata – Notochord at some stage and separate anal and urinary openings to the outside

Explanation:

Mollusca mostly oviparous and a few viviparous. The development may be direct or indirect with trochophore, velliger, and glochidium.

Question 34.

Which one of the following groups of three animals each is correctly matched with their one characteristic morphological feature?

[4 Marks]

(A)

Animals		Morphological feature
Liver fluke, Sea anemone, Sea	-	Bilateral symmetry

(B)

Animals		Morphological feature
Centipede, Prawn Sea urchin	-	Jointed appendages

(C)

Animals		Morphological feature
Scorpion, Spider, Cockroach	-	Ventral solid central nervous system

(D)

Animals		Morphological feature
Cockroach, Locust, Taenia	-	Metameric segmentation

Explanation:

Scorpion, spider, and cockroach have a ventral solid central nervous system. All three belong to phylum Arthropoda.

— Spider belongs to class Arachnida

— Scorpion belongs to class Arachnida

— Cockroach belongs to class Insecta

Question 35.

Which one of the following is the true description about an animal concerned?

[4 Marks]

(A) Cockroach – 10 pairs of spiracles (2 pairs on thorax and 8 pairs on abdomen)

(B) Rat – Left kidney is slightly higher in position than the right one

(C) Frog – Body divisible into three regions – head, neck and trunk

(D) Earthworm – The alimentary canal consists of a sequence of pharynx, oesophagus, stomach, gizzard and intestine

Explanation:

Cockroach belongs to phylum Arthropoda. It has 10 pairs of Spiracles (2 pairs on thorax and 8 pairs on abdomen).

Question 36.

Replum is present in the ovary of flower of

[4 Marks]

(A) Pea

(B) Sun flower

(C) Lemon

(D) Mustard

Explanation:

Replum is present in the ovary of flower of mustard. A false septun called replum develops between the two parietal placentae in mustard and other members of family

Brassicaceae. The ovary becomes bilocular.

Question 37.

The length of different internodes in a culm of sugarcane is variable because of

[4 Marks]

(A) shoot apical meristem

(B) position of axillary buds

(C) intercalary meristem

(D) size of leaf lamina at the node below each internode

Explanation:

The length of different internodes in a culm of sugarcane is variable because of intercalary meristem. Intercalary meristem is not a part of apical meristem. It occurs in the internodes of grasses (sugarcane) between leaf nodes and enables longitudinal growth of the stem.

Question 38.

Vascular tissues in flowering plants develop from

[4 Marks]

(A) phellogen

(B) plerome

(C) periblem

(D) dermatogen

Explanation:

Vascular tissues in flowering plants develop from plerome. Plerome is a central core of primary meristem which gives rise to all cells of the stele from the pericycle inward.

Question 39.

Which one of the following is resistant to enzyme action?

[4 Marks]

(A) Cork

(B) Wood fibre

(C) Leaf cuticle

(D) Pollen exine

Explanation:

Pollen exine is resistant to enzyme action. The hard outer layer called the exine is made up of sporopollenin which is one of the most resistant organic material known.

It can withstand high temperatures and strong acids and alkali. No enzymes that degrade sporopollenin are so far known.

Question 40.

Earthworms have no skeleton but during burrowing, the anterior end becomes turgid and acts as a hydraulic skeleton. It is due to

[4 Marks]

(A) setae

(B) gut peristalsis

(C) coelomic fluid

(D) blood

Explanation:

Earthworms have no skeleton but during burrowing, anterior end becomes turgid and acts as a hydraulic skeleton. It is due to coelomic fluid. The coelom is filled with an alkaline milky coelomic fluid.

Question 41.

Keeping in view the fluid mosaic model for the structure of cell membrane, which one of the following statements is correct with respect to movement of lipids and proteins from one lipid monolayer to the other (described as flip-flop movement)?

[4 Marks]

(A) Both lipids and proteins can flip-flop

(B) While lipids can rarely flip-flop, protein can not

(C) Neither lipids, nor proteins can flip-flop

(D) While proteins can flip-flop, lipids can not

Explanation:

Keeping in view of the ‘Fluid mosaic model’ for the structure of cell membrane, the movement of lipids and proteins from one lipid monolayer to the other, lipid can rarely flip-flop, protein can not. The fluid mosaic model is the most important and widely accepted latest model for plasma membrane was given by Singer and Nicolson in 1972. According to them it resembles “protein icebergs in a sea of lipids”.

Question 42.

Cellulose is the major component of cell walls of

[4 Marks]

(A) Pseudomonas

(B) Saccharomyces

(C) Xanthomonas

(D) Pyhtium

Explanation:

Cellulose is a major component of cell wall of Pythium. Pythium is a genus of parasitic Oomycete. Because this group of organisms were once classified as fungi, they are sometimes still treated as such. Pythium, like others in the family Pythiaceae, are usually characterized by their production of coenocytic hyphae, hyphae without separations. Oogonia generally contain a single oospore. Antheridia contain an elongated and club shaped antheridium.

Question 43.

The two sub-units of ribosome remain united at critical ion level of

[4 Marks]

(A) magnesium

(B) calcium

(C) copper

(D) manganese

Explanation:

The two subunits of ribosome remain united at a critical ion level of magnesium. The presence of magnesium and its amount plays an important role in the appearance and structure of the ribosomes.

Question 44.

Vacuole in a plant cell

[4 Marks]

(A) lacks membrane and contains air

(B) lacks membrane and contains water and excretory substances

(C) is membrane-bound and contains water and excretory substances

(D) is membrane-bound and contains storage proteins and lipids

Explanation:

Vacuole in a plant cell is membrane bound and contains water and excretory substance. Vacuole is a space within the cytoplasm of living cell that is filled with air,

water or other liquid, cell sap or food particles. In plant cells, there is usually one large vacuole bounded by a single layered membrane (tonoplast or vacuole membrane);

animal cells usually have several small vacuoles.

Question 45.

A competitive inhibitor of succinic dehydrogenase is

[4 Marks]

(A) m-ketoglutarate

(B) Malate

(C) Oxaloacetate

(D) Malonate

Explanation:

Malonate is a powerful inhibitor of cellular respiration, because it binds to the active site of succinate dehydrogenase in the citric acid cycle but does not react. It thereby competes with succinate, since it does not have the $-CH_2,-CH_2,-$ group (as in succinate which is required for dehydrogenation). For the oxidative phosphorylation reaction, malonate is an inhibitor for complex II which again contains succinate dehydrogenase.

Question 46.

The rupture and fractionation do not usually occur in the water column in vessel/tracheids during the ascent of sap because of

[4 Marks]

(A) weak gravitational pull

(B) transpiration pull

(C) cohesion and adhesion

(D) lignified thick walls

Explanation:

Rupture and fractionation do not usually occur in the water column in vessel/ tracheids during the ascent of sap because of cohesion and tension.

Question 47.

Carbohydrates are commonly found as starch in plant storage organs. Which of the following five properties of starch (A–E) make it useful as a storage material?

(A) Easily translocated

(B) Chemically non-reactive

(C) Easily digested by animals

(D) Osmotically inactive

(E) Synthesized during photosynthesis

The useful properties are

[4 Marks]

(A) (A) and (E)

(B) (A), (C) and (E)

(C) (B) and (D)

(D) (B) and (C)

Explanation:

Carbohydrates are commonly found as starch in plant storage organs. It is a chemically non-reactive and osmotically inactive polysaccharide with a molecular weight. Carbohydrates perform a vital role in living organisms. Starch and other polysaccharides serve as energy storage in plants, particularly in seeds, tubers, etc. which provide a major energy source for animals, including humans.

Question 48.

Nitrogen fixation in root nodules of Alnus is brought about by:

[4 Marks]

(A) Azorhizobium

(B) Frankia

(C) Clostridium

(D) Bradyrhizobium

Explanation:

Nitrogen fixation in root nodules of *Alnus* is brought about by *Frankia*. It is a symbiotic nitrogen-fixing bacteria in root nodules of several non-legume plants like Casuarina (Australian Pine), Myrica and Alnus (Alder).

Question 49.

Electrons from excited chlorophyll molecule of photosystem II are accepted first by:

[4 Marks]

(A) Cytochrome-f

(B) Cytochrome-b

(C) Quinone

(D) Ferredoxin

Explanation:

Electrons from excited chlorophyll molecule of photosystem II are accepted first by quinone. Photosystem II is a photosynthetic pigment system along with some electron carriers that are located in the appressed part of the grana thylakoids. Photosystem II has chlorophyll a, b and carotenoids. Other components of PS II are pheophytin, plastoquinone (PQ), cytochrome complex and blue coloured copper containing plastocyanin.

Question 50.

In leaves of C_4 plants malic acid synthesis during CO_2 fixation occurs in

[4 Marks]

(A) guard cells

(B) bundle sheath

(C) mesophyll cells

(D) epidermal cells

Explanation:

In leaves of C_4 plants, malic acid synthesis during CO_2 fixation occurs in mesophyll cells. The mesophyll cells perform C_4 cycle and the cells of bundle sheath perform C_3 cycle.

Question 51.

The energy-releasing process in which the substrate is oxidised without an external electron acceptor is called

[4 Marks]

(A) glycolysis

(B) aerobic respiration

(C) photorespiration

(D) fermentation

Explanation:

Glycolysis takes place in all body cells and is of two types—

(a) Anaerobic glycolysis—From glycogen or glucose to lactic acid in muscles.

(b) Aerobic glycolysis—From glycogen or glucose to pyruvic acid (all cells of body).

Question 52.

The chemiosmotic coupling hypothesis of oxidative phosphorylation proposes that adenosine triphosphate (ATP) is formed because:

[4 Marks]

(A) there is a change in the permeability of the inner mitochondrial membrane toward adenosine diphosphate (ADP)

(B) a proton gradient forms across the inner membrane

(C) ADP is pumped out of the matrix into the intermembrane space

(D) high energy bonds are formed in mitochondrial proteins

Explanation:

Chemiosmotic theory was postulated by the British biochemist Peter Mitchell (1920–22) to explain the formation of ATP in the mitochondrial electron transport chain. As electrons are transferred along the electron carrier system in the inner mitochondrial membrane, hydrogen ions (protons) are actively transported into the space between the inner and outer mitochondrial membranes, which thus contains a higher concentration of protons than the matrix. This creates an electrochemical gradient across the inner membrane, through which protons move back into the matrix.

Question 53.

In germinating seeds fatty acids are degraded exclusively in the

[4 Marks]

(A) peroxisomes

(B) mitochondria

(C) proplastids

(D) glyoxysomes

Explanation:

In germinating seeds, fatty acids are degraded exclusively in the glyoxysomes.

Glyoxylate cycle occurs in tissues rich in fats, such as those of germinating seeds; the enzymes involved in the cycle, which have not been found in mammals, are contained in organelles called glyoxysomes.

Question 54.

Senescence as an active developmental cellular process in the growth and functioning of a flowering plant, is indicated in

[4 Marks]

(A) floral parts

(B) leaf abscission

(C) annual plants

(D) vessels and tracheid differentiation

Explanation:

Senescence is an active developmental cellular process in the growth and functioning of a flowering plant, indicated in leaf abscission. Senescence is the changes that occur in an organism between maturity and death, i.e. ageing.

Characteristically there is a deterioration in functioning as the cells become less efficient in maintaining and replacing vital cell components.

Question 55.

Importance of day length in flowering of plants was first shown in

[4 Marks]

(A) lemma

(B) tobacco

(C) cotton

(D) petunia

Explanation:

Photoperiodism was first studied by Garner and Allard (1920). They observed that “Maryland Mammoth” variety of tobacco could be made to flower in summer by reducing the light hours with artificial darkening. It could be made to remain vegetative in winter by providing extra light.

Question 56.

Which one of the following is the correct matching of the site of action on the given substrate, the enzyme acting upon it and the end product?

[4 Marks]

(A) Stomach: Fats Micelles

(B) Small intestine: Proteins Amino acids

(C) Duodenum: Triglycerides Monoglycerides

(D) Small intestine: Starch Disaccharide (Maltose)

Explanation:

Small intestine- Starch disaccharide (maltose)

Small intestine is the portion of the alimentary canal between the stomach and large intestine. It is subdivided into the duodenum, jejunum and ileum. It plays an essential

role in the final digestion and absorption of food.

Question 57.

What will happen if the secretion of parietal cells of gastric glands is blocked with an inhibitor?

[4 Marks]

(A) Enterokinase will not be released from the duodenal mucosa and so trypsinogen is not converted to trypsin

(B) In the absence of HCl secretion, inactive pepsinogen is not converted into the active enzyme pepsin

(C) Gastric juice will be deficient in pepsinogen

(D) Gastric juice will be deficient in chymosin

Explanation:

If the secretion of parietal cells of gastric glands is blocked with an inhibitor, in the absence of HCI secretion, inactive pepsinogen is not converted into the active enzyme pepsin.

Question 58.

Which type of white blood cells are concerned with the release of histamine and the natural anti-coagulant heparin?

[4 Marks]

(A) Monocytes

(B) Eosinophils

(C) Neutrophils

(D) Basophils

Explanation:

Basophil is a type of white blood cell (leucocyte) that has a lobed nucleus surrounded by granular cytoplasm. Basophils are produced continually by stem cells in the red bone marrow & move about in an amoeboid fashion. Like mast cells, they produce histamine and heparin as part of the body's defences at the site of an infection or injury.

Question 59.

The most active phagocytic white blood cells are:

[4 Marks]

(A) eosinophils and lymphocytes

(B) neutrophils and monocytes

(C) neutrophils and eosinophils

(D) lymphocytes and macrophages

Explanation:

The most active phagocytic white blood cells are neutrophils and monocytes which modulate the innate immune response and phagocytose infectious agents.

Neutrophil has a lobed nucleus and granular cytoplasm. Neutrophils engulf bacteria and release various substances such as lysozyme and oxidising agents.

Monocyte is the largest form of WBC in vertebrates. It has a kidney-shaped nucleus and is actively phagocytic, ingesting bacteria and cell debris.

Question 60.

The haemoglobin of a human foetus

[4 Marks]

(A) has a lower affinity for oxygen than that of the adult.

(B) has affinity for oxygen the same as that of an adult.

(C) has only 2 protein subunits instead of 4.

(D) has a higher affinity for oxygen than that of an adult.

Explanation:

The haemoglobin of a human foetus has a higher affinity for oxygen than that of an adult. Foetal haemoglobin binds to oxygen more strongly than adult haemoglobin, enabling the transfer of oxygen from mother to foetus prenatally.

Question 61.

In humans, blood passes from the post caval to the diastolic right atrium of heart due to

[4 Marks]

(A) suction pull

(B) pushing open of the venous valves

(C) stimulation of the sino auricular node

(D) pressure difference between the post caval and atrium

Explanation:

In humans, blood passes from the post caval to the diastolic right atrium of heart due to suction pull.

Question 62.

Which one of the following items gives its correct total number?

[4 Marks]

(A) Cervical vertebrae in humans – 8

(B) Types of diabetes – 3

(C) Floating ribs in humans – 4

(D) Amino acids found in proteins – 16

Explanation:

In humans total number of ribs are Fourteen (14). These are of three types:

(1) True ribs – These are 7 in number.

(2) False ribs – These are 2 in number.

(3) Floating ribs – These are 4 in number.

Question 63.

During the propagation of a nerve impulse, the action potential results from the movement of:

[4 Marks]

(A) Na^+ ions from extracellular fluid to intracellular fluid

(B) Na^+ ions from intracellular fluid to extracellular fluid

(C) K^+ ions from extracellular fluid to intracellular fluid

(D) K^+ ions from intracellular fluid to extracellular fluid

Explanation:

Total sum of physio-electrochemical changes that takes place along the length of nerve fibre is known as nerve impulse.

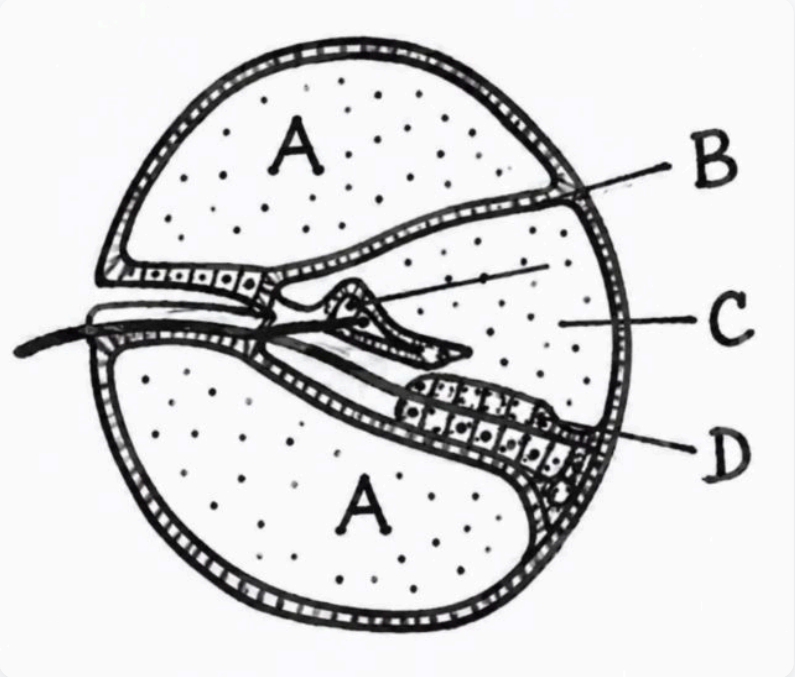
Change in potential due to stimulation of nerve fibre is called action potential. During propagation of nerve impulse, Na^+ enters inside so (+ve) charge is formed inside the membrane. K^+ ions come out.

Question 64.

Given below is a diagrammatic cross section of a single loop of human cochlea:

Which one of the following options correctly represents the names of three different parts?

[4 Marks]



(A) D: Sensory hair cells, A: Endolymph, B: Tectorial membrane

(B) A: Perilymph, B: Tectorial membrane, C: Endolymph

(C) B: Tectorial membrane, C: Perilymph, D: Secretory cells

(D) C: Endolymph, D: Sensory hair cells, A: Serum

Explanation:

The diagrammatic cross section of a single loop of human cochlea represents three different parts—

A - Perilymph

B - Tectorial membrane

C - Endolymph

Cochlea arises from sacculus. It is spirally coiled duct. It is also known as Lagena.

It is connected with sacculus by duct of reuniens.

Question 65.

Which one of the following is the correct difference between Rod Cells and Cone Cells of our retina?

[4 Marks]

(A)

	Rod Cells	Cone Cells
Visual: acuity	High	Low

(B)

	Rod Cells	Cone Cells
Visual pigment contained :	Iodops	in Rhodpsin

(C)

	Rod Cells	Cone Cells
Overall function :	Vision in poor light	Colour vision and detailed vision in bright light

(D)

	Rod Cells	Cone Cells
Distribution :	More concentrated in retina	Evenly distributed all over

Explanation:

Rod cells is a type of light sensitive receptor cell present in the retinas of vertebrates. Rod contains the pigment rhodopsin and are essential for vision in dim light.

Cone cells is a type of light sensitive receptor cell, found in retinas of all diurnal vertebrates. Cones are specialized to transmit information about colour and are responsible for the visual acuity of the eye. They function best in bright light.

Question 66.

A cell can be balanced against 110 cm and 100 cm of potentiometer wire, respectively with and without being short circuited through a resistance of 10Q. Its internal resistance is

[4 Marks]

(A) 2.0 ohm

(B) zero

(C) 0.5 ohm

(D) 1.0 ohm

Explanation:

Here E > ER / R + r, hence the lengths 110 cm and 100 cm are interchanged.

Without being short-circuited through R, only the battery E is balanced.

E = V / L × l_1 = V / L ×110 ...(i)

When R is connected across E, Ri = V / L× l_2

or, R (E / R+r) = V / L × 100

Diving (i) by (ii), we get

R + r / R = 110 / 100

or, 100 R +100r = 110 R

or, 10R = 100 r

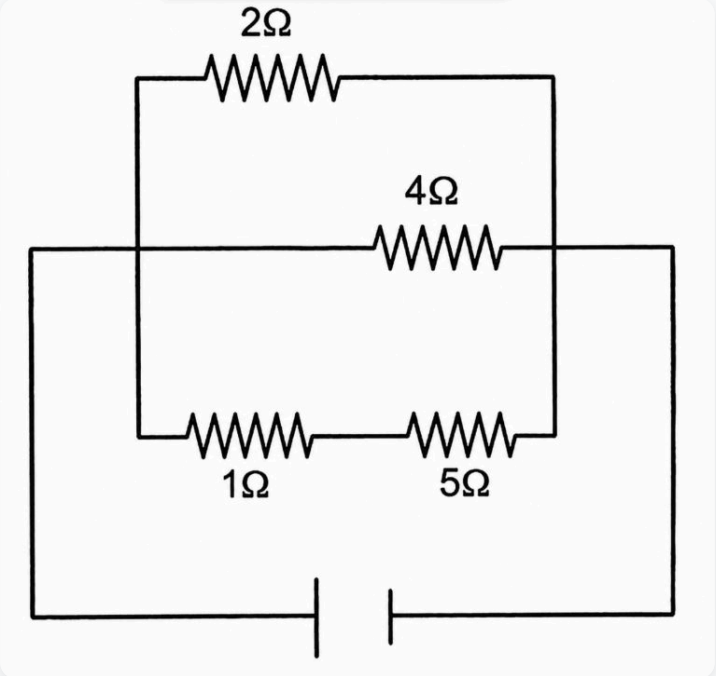
∴ r = 10R / 100 = ((10 × 10) / 100) (∴R = 10Ω)

⇒ r = 1 Ω

Question 67.

A current of 3 amp flows through the 2Ω resistor shown in the circuit. The power dissipated in the 5-Ω resistor is:

[4 Marks]



(A) 1 watt

(B) 5 watt

(C) 4 watt

(D) 2 watt

Explanation:

Clearly, 2Ω, 4Ω and (1+5) Ω resistors are in parallel.

Hence, potential difference is same across each of them.

$\therefore I_1 \times 2 = I_2 \times 4 = I_3 \times 6$

Given $I_1 = 3A$ ($\therefore I_1 \times 2 = I_3 \times 6$)

Given $I_1 = 3A$

$\therefore I_1 \times 2 = I_3 \times 6$ provides

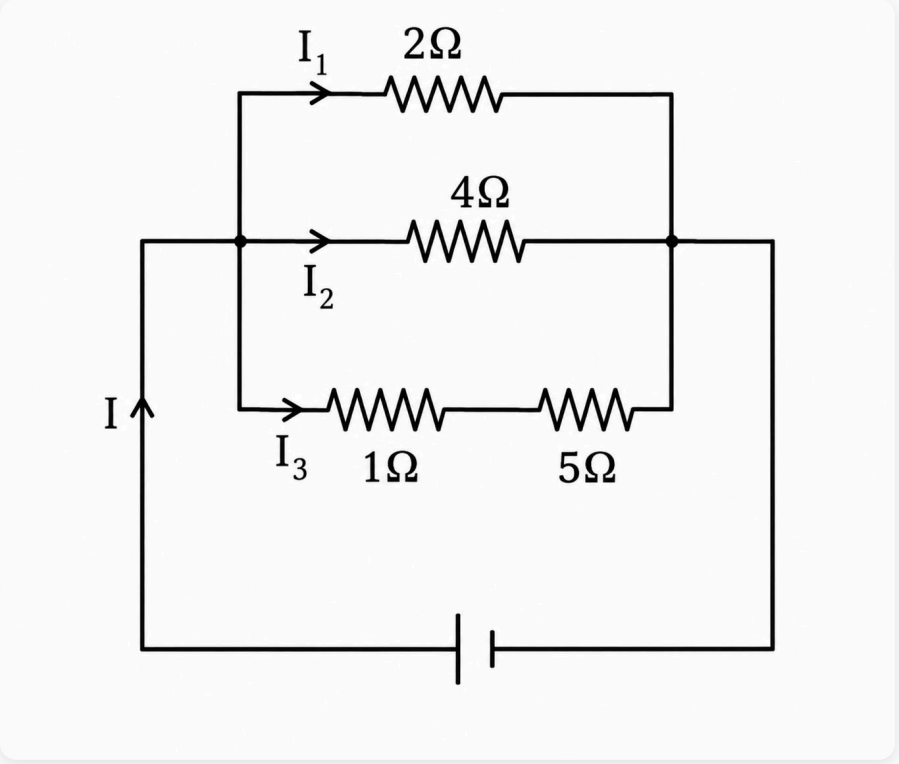
$I_3 = (I_1 \times 2 / 6) = (3 \times 2 / 6) = 1A.$

Now, the potential across the 5Ω resistor is

$V = I_3 \times 5 = 1 \times 5 = 5V$

the power dissipated in the 5Ω resistor

$P = V^2 / R = 5^2 / 5 = 5watt.$



Question 68.

A particle of mass m, charge Q and kinetic energy T enters a transverse uniform magnetic field of induction □□. After 3 seconds, the kinetic energy of the particle will be:

[4 Marks]

(A) 3T

(B) 2T

(C) T

(D) 4T

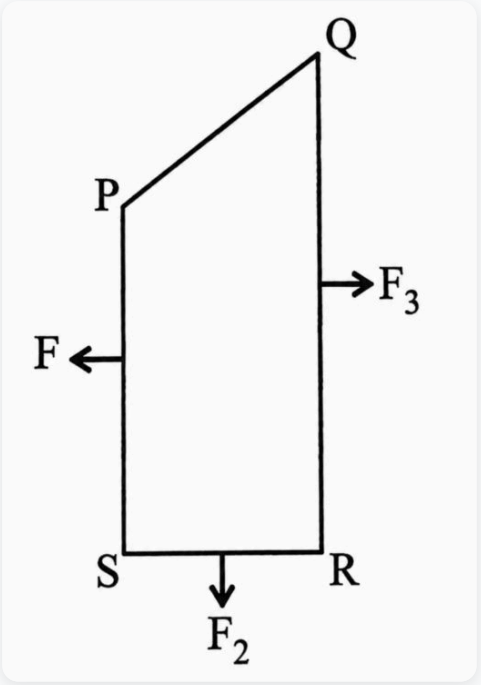
Explanation:

When a charged particle enters a transverse magnetic field it traverse a circular path. Its kinetic energy remains constant.

Question 69.

A closed loop PQRS carrying a current is placed in a uniform magnetic field. If the magnetic forces on segments PS, SR, and RQ are F_1, F_2, and F_3 respectively and are in the plane of the paper and along the directions shown, the force on the segment QP is

[4 Marks]



(A) $F_3 - F_1 - F_2$

(B) $\sqrt{(F_3 - F_1)^2 + F_2^2}$

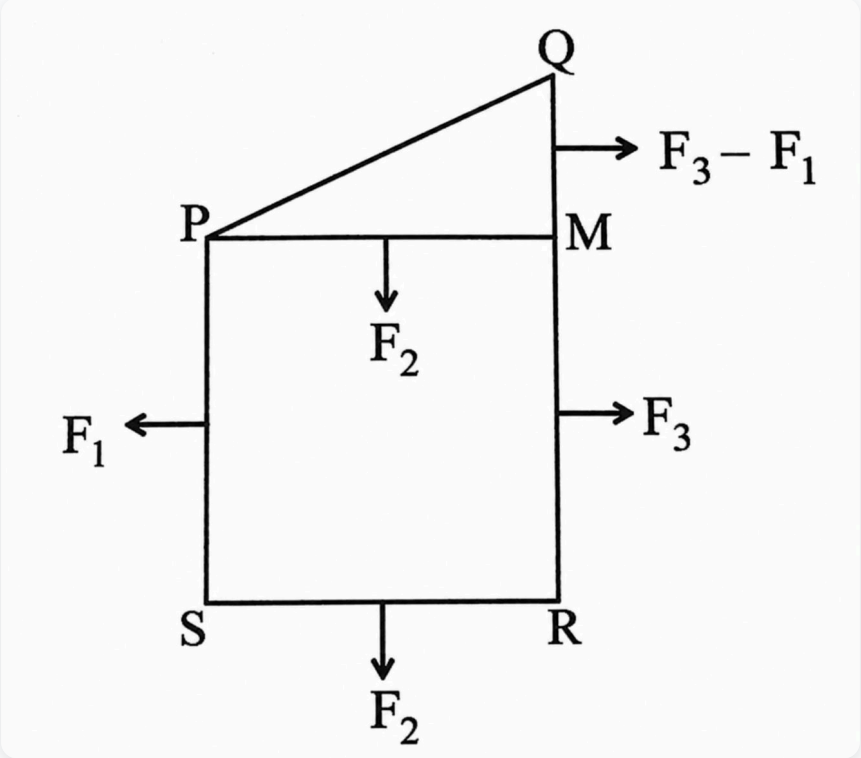
(C) $\sqrt{(F_3 - F_1)^2 - F_2^2}$

(D) $F_3 - F_1 + F_2$

Explanation:

According to the figure the magnitude of force on the segment QM is F_3 —F_1, and PM is F_2.

Therefore, the magnitude of the force on segment PQ is $\sqrt{(F_3-F_1)^2 + F_2^2}$



Question 70.

A galvanometer of resistance 50 Ω is connected to battery of 3V along with a resistance of 2950 Ω in series. A full scale deflection of 30 divisions is obtained in the galvanometer. In order to reduce this deflection to 20 divisions, the resistance in series should be

[4 Marks]

(A) 4450 Ω

(B) 5550 Ω

(C) 6050 Ω

(D) 5050 Ω

Explanation:

Total internal resistance = (50 +2950)Ω = 3000Ω

Emf of the cell, ε= 3V

∴ Current = ε / R = 3 / 3000 = 1×10^{-3}A = 1.0mA

∴ Current for full scale deflection of 30 division is 1.0mA

∴ Current for a deflection of 20 divisions,

I = (20 / 30 ×1) mA or I = (2 / 3)mA

Let the resistance be x Ω. Then

$x = \varepsilon / I = (3V / (2/3 \times 10^{-3} A)) = (3 \times 3 \times 10^3 / 2)\ \Omega = 4500\Omega$

But the resistance of the galvanometer is 50Ω.

∴ Resistance to be added

$= (4500-50)\Omega = 4450\Omega$

Question 71.

Curie temperature is the temperature above which

[4 Marks]

(A) ferromagnetic material becomes paramagnetic material

(B) paramagnetic material becomes diamagnetic material

(C) paramagnetic material becomes ferromagnetic material

(D) ferromagnetic material becomes diamagnetic material

Explanation:

Curie temperature is the temperature above which ferromagnetic material becomes paramagnetic material.

Question 72.

A circular disc of radius 0.2 meter is placed in a uniform magnetic field of induction $1/\pi$ (Wb/m^2) in such a way that its axis makes an angle of 60° with □□. The magnetic flux linked with the disc is:

[4 Marks]

(A) 0.01 Wb

(B) 0.08 Wb

(C) 0.06 Wb

(D) 0.02 Wb

Explanation:

Here, $B = 1 / \pi$ (Wb/m^2)

$\theta = 60°$

Area normal to the plane of the disc

$=\pi r^2 \cos 60° = \pi r^2 / 2$

Flux = B × normal area

$= 0.2 \times 0.2 / 2 = 0.02Wb$

Question 73.

A long solenoid has 500 turns. When a current of 2 ampere is passed through it, the resulting magnetic flux linked with each turn of the solenoid is 4×10^{-3} Wb. The self-inductance of the solenoid is

[4 Marks]

(A) 40 henry

(B) 1.0 henry

(C) 2.0 henry

(D) 2.5 henry

Explanation:

Total number of turns in the solenoid, N= 500

Current , I = 2A.

Magnetic flux linked with each turn

$= 4\times10^{-3}Wb$

As, $\Phi = LI$ or $N\Phi = LI \Rightarrow L = N\Phi / 1$

$= (500 \times 4 \times 10^{-3} / 2)$ henry = 1 H.

Question 74.

In an a.c. circuit the e.m.f. (e) and the current (i) at any instant are given respectively by

$e = E_0 \sin \omega t$

$i = I_0 \sin (\omega t - \phi)$

The average power in the circuit over one cycle of a.c. is

[4 Marks]

(A) $E_0 I_0$

(B) $(E_0I_0 / 2)\sin \Phi$

(C) $(E_0I_0 / 2)\cos \Phi$

(D) $E_0I_0 / 2$

Explanation:

The average power in the circuit over one cycle of a.c. is given by

$P_{av} = e_{rms} \times i_{rms} \times \cos\Phi$

$= E_0 / \sqrt{2} \times I_0 / \sqrt{2} \times \cos \Phi = (E_0 I_0 / 2) \cos\Phi$

Question 75.

The velocity of electromagnetic radiation in a medium of permittivity ε_0 and permeability μ_0 is given by

[4 Marks]

(A) $\sqrt{(\mu_0 \varepsilon_0)}$

(B) $\sqrt{(\varepsilon_0 / \mu_0)}$

(C) $1 / \sqrt{(\mu_0 \varepsilon_0)}$

(D) $\sqrt{(\mu_0 / \varepsilon_0)}$

Explanation:

The velocity of electromagnetic radiation in a medium of permittivity ε_0 and permeability μ_0 is equal to $1/\sqrt{(\mu_0 \varepsilon_0)}$.

Question 76.

Two thin lenses of focal lengths f_1 and f_2 are in contact and coaxial. The power of the combination is:

[4 Marks]

(A) $f_1 + f_2 / f_1f_2$

(B) $f_1 + f_2 / 2$

(C) $\sqrt{(f_1 / f_2)}$

(D) $\sqrt{(f_2 / f_1)}$

Explanation:

The focal length of the combination

$1 / f = 1 / f_1 + 1 / f_2$

∴ Power of the combinations,

$P = f_1 + f_2 / f_1 f_2$ (∵ $P = 1 / f$)

Question 77.

A boy is trying to start a fire by focusing sunlight on a piece of paper using an equiconvex lens of focal length 10 cm. The diameter of the Sun is 1.39×10^9 m and its mean distance from the earth is 1.5×10^{11} m. What is the diameter of the Sun's image on the paper?

[4 Marks]

(A) 12.4×10^{-4}m

(B) 6.5×10^{-5}m

(C) 6.5×10^{-4}m

(D) 9.2×10^{-4}m

Explanation:

We have, $|v / u|$ = size of image / size of object

or, size of image = $|v / u| \times$ size of object

$= (10^{-1} / 1.5 \times 10^{11}) \times (1.39 \times 10^9)$

$= 0.92 \times 10^{-3}$ m $= 9.2 \times 10^{-4}$m

∴ Diameter of the sun's image = 9.2×10^{-4}m

Question 78.

The periodic waves of intensities I_1 and I_2 pass through a region at the same time in the same direction. The sum of the maximum and minimum intensities is:

[4 Marks]

(A) $(\sqrt{I_1} - \sqrt{I_2})^2$

(B) $(\sqrt{I_1} + \sqrt{I_2})^2$

(C) $I_1 + I_2$

(D) $2(I_1 + I_2)$

Explanation:

The resultant intensity of two periodic waves at a point is given by

$I = I_1 + I_2 + 2\sqrt{(I_1 I_2)} \cdot \cos\phi$.

Resultant intensity is maximum if

$\cos\phi = -1$

i.e. $I_{max} = I_1 + I_2 + 2\sqrt{(I_1 I_2)}$

Resultant intensity is minimum if

$\cos\phi = +1$

i.e., $I_{min} = I_1 + I_2 - 2\sqrt{(I_1 I_2)}$

Therefore, the sum of the maximum and minimum intensities is

$I_{max} + I_{min}$

$= I_1 + I_2 + 2\sqrt{(I_1 I_2)} + I_1 + I_2 - 2\sqrt{(I_1 I_2)}$

$= 2(I_1 + I_2)$

Question 79.

In the phenomenon of electric discharge through gases at low pressure, the coloured glow in the tube appears as a result of:

[4 Marks]

(A) collisions between the charged particles emitted from the cathode and the atoms of the gas

(B) collision between different electrons of the atoms of the gas

(C) collision between the atoms of the gas

(D) excitation of electrons in the atoms

Explanation:

In discharge tube, collisions between the charged particles emitted from the cathode and atoms of the gas results to the coloured glow in the tube.

Question 80.

A particle of mass 1 mg has the same wavelength as an electron moving with a velocity of 3×10^6 ms^{-1}. The velocity of the particle is:

[4 Marks]

(A) $9 \times 10^{-2}\ ms^{-1}$

(B) $3 \times 10^{-31}\ ms^{-1}$

(C) $2.7 \times 10^{-21}\ ms^{-1}$

(D) $2.7 \times 10^{-18}\ ms^{-1}$

Explanation:

Wavelength of particle

$(\lambda_1) = h / mv = (h / (1 \times 10^{-6}) \times v)$

where v is the velocity of the particle.

Wave length of electron,

$(\lambda_2) = (h / (9.1 \times 10^{-31}) \times (3 \times 10^6))$

but $\lambda_1 = \lambda_2$ (Given)

$\therefore (h / (1 \times 10^{-6}) \times v) = (h / (9.1 \times 10^{-31}) \times (3 \times 10^6))$

$\Rightarrow v = 9.1 \times 10^{-31} \times 3 \times 10^6 / 10^{-6} = 2.73 \times 10^{-18}\ ms^{-1}$

Question 81.

The work function of a surface of a photosensitive material is 6.2 eV. The wavelength of incident radiation for which the stopping potential is 5 V lies in the:

[4 Marks]

(A) Ultraviolet region

(B) Visible region

(C) Infrared region

(D) X-ray region

Explanation:

Work function, $\Phi_0 = 6.2eV$.

Stopping Potential $V_0 = 5V$.

As, $eV_0 = hv - \Phi_0$

or $eV_0 + \Phi_0 = hc / \lambda$

or $\lambda = (hc / eV_0 + \Phi_0)$

$\lambda = (6.62 \times 10^{-34}\ Js \times 3 \times 10^8 ms^{-1} / 1.6 \times 10^{-19} \times 5J + 6.2 \times 1.6 \times 10^{-19}J)$

$= (6.62 \times 10^{-34} \times 3 \times 10^8 / 1.6 \times 10^{-19} (5+6.2))\ m$

$= (19.86 \times 10^{-26} / 17.92 \times 10^{-19}) = 1.10 \times 10^{-7}\ m.$

Question 82.

The ground state energy of hydrogen atom is 13.6 eV. When its electron is in the first excited state, its excitation energy is

[4 Marks]

(A) 0

(B) 10.2 eV

(C) 6.8 eV

(D) 3.4 eV

Explanation:

When the electron is in first excited state

(n = 2), the excitation energy is given by

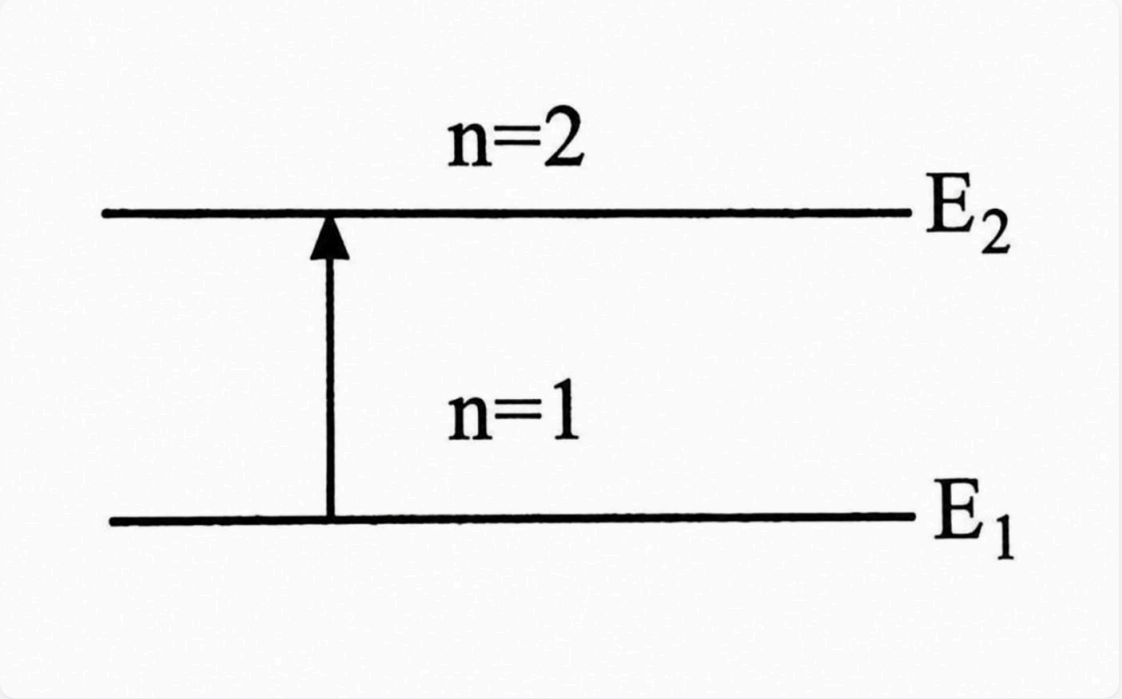
$\Delta E = E_2 - E_1$

We have, $E_n = (-13.6 / n^2)$ eV

$\therefore E_2 = (-13.6 / 2^2)$ eV = -3.4eV

Given E_1 = -13.6eV

$\therefore \Delta E = (-3.4) - (-13.6) = 10.2$ eV.



Question 83.

Two nuclei have their mass numbers in the ratio of 1 : 3. The ratio of their nuclear densities would be

[4 Marks]

(A) 1 : 3

(B) 3 : 1

(C) $(3)^{1/3}$: 1

(D) 1 : 1

Explanation:

Required ratio of nuclear densities = ρ_1 / ρ_2

= $((M_1/V_1) / (M_2/V_2)) = M_1/M_2 \times V_2/V_1 = 1 / 3 \times ((4/3)\pi R_2^3 / (4/3)\pi R_1^3$

= $1 / 3 \times (R_2 / R_1)^3 = 1 / 3 \times (R_0 M_2^{1/3} / R_0 M_1^{1/3})^3$ [$\because R = R_0 M^{1/3}$]

= $1 / 3 \times (M_2 / M_1) = 1 / 3 \times (3 / 1)$

= 1 : 1

Question 84.

If M(A; Z), M_p, and M_n denote the masses of the nucleus ${}^{A}_ZX$, proton and neutron respectively in units of u (1 u = 931.5 MeV/c^2) and BE represents its bonding energy in MeV, then

[4 Marks]

(A) $M(A, Z) = ZM_p + (A - Z)M_n - BE/c^2$

(B) $M(A, Z) = ZM_p + (A - Z)M_n + BE$

(C) $M(A, Z) = ZM_p + (A - Z)M_n + BE/c^2$

(D) $M(A, Z) = ZM_p + (A - Z)M_n - BE$

Explanation:

Mass defect = $ZM_p + (A - Z)M_n - M(A, Z)$

or, $B.E. / c^2 = ZM_p + (A - Z)M_n - M(A, Z)$

$\therefore M(A, Z) = ZM_p + (A - Z)M_n - B.E. / c^2$

Question 85.

Two radioactive materials X_1 and X_2 have decay constants 5λ and λ respectively. If initially they have the same number of nuclei, then the ratio of the number of nuclei of X_1 to that of X_2 will be 1 / e after a time

[4 Marks]

(A) e / λ

(B) 1 / 4λ

(C) λ

(D) (1 / 2) λ

Explanation:

Let the required time be t. Then

$N_1 = N_0e_(-\lambda_1 t)$; $N_2 = N_0e_(-\lambda_2 t)$

Where

N_1 = number of nuclei of X_1 after time t

N_2 = number of nuclei of X_2 after time t

N_0 = initial number of nuclei of X_1 and X_2 each

Now, $N_1 / N_2 = ((N_0e_(-\lambda_1 t) / N_0e_(-\lambda_2 t))$ Here $N_1 / N_2 = 1 / e$

$\lambda_1 = 5\lambda$; $\lambda_2 = \lambda$

∴ 1 / e = (e_(−5λt) / e_(−λt)) ⇒ e^{-1} = e_(−4λt) ⇒ 4λt = 1

∴ t = 1 / 4λ

Question 86.

A p-n photodiode is made of a material with a band gap of 2.0 eV. The minimum frequency of the radiation that can be absorbed by the material is nearly

[4 Marks]

(A) 20 × 10^{14} Hz

(B) 1 × 10^{14} Hz

(C) 5 × 10^{14} Hz

(D) 10 × 10^{14} Hz

Explanation:

E_g = 2.0 eV = 2 × 1.6 × 10^{-19} J

∴ E_g = hν

∴ ν = E_g / h

= ((2 × 1.6 × 10^{-19} J) / (6.62 × 10^{-34}Js))

= 0.4833 × 10^{15} s^{-1} = 4.833 × 10^{14} Hz ≃ 5×10^{14} Hz

Question 87.

The voltage gain of an amplifier with 9% negative feedback is 10. The voltage gain without feedback will be

[4 Marks]

(A) 90

(B) 10

(C) 1.25

(D) 100

Explanation:

Negative feedback is applied to reduce the output voltage of an amplifier. If there is no negative feedback, the value of output voltage could be very high. In the options given, the maximum value of

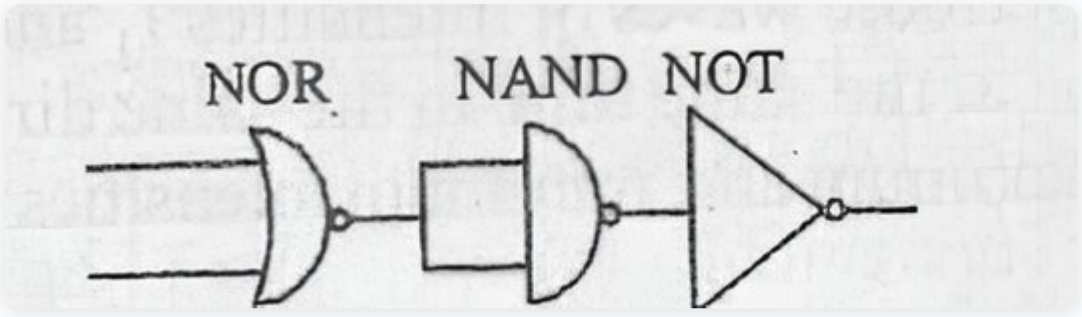
voltage gain is 100. Hence it is the correct option.

Question 88.

The circuit

is equivalent to

[4 Marks]



(A) AND gate

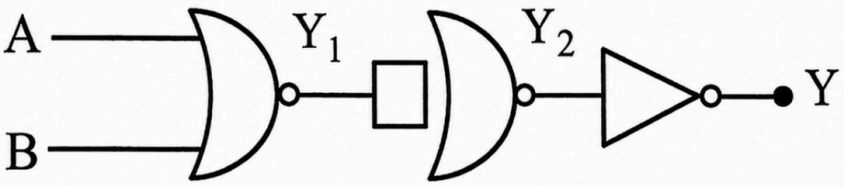
(B) NAND gate

(C) OR gate

(D) NOR gate

Explanation:

Let A and B be inputs and Y the output.



Then $Y_1 = \left(\overline{A + B}\right) = \overline{A}.\overline{B}$

(By De-Morgan's theorem)

$Y_2 = \overline{Y_1} = \left(\overline{\overline{A}.\overline{B}}\right) = A + B = A + B$

$\therefore\ Y = \overline{Y_2} = \left(\overline{A + B}\right)$

Hence, the given circuit is equivalent to a NOR gate.

Question 89.

If the lattice parameter for a crystal line structure is 3.6 Å, then the atomic radius in fcc crystal is

[4 Marks]

(A) 1.27 Å

(B) 2.92 Å

(C) 2.10 Å

(D) 1.81 Å

Explanation:

Atomic radius is the distance between two nearest neighbours. For fcc crystal

Atomic radius = $(1 / 2\sqrt{2}) \times$ lattice parameter

or $r = (a / (2\sqrt{2}))$

$r = (3.6 / (2\sqrt{2}))$ Å = 1.27 Å

Question 90.

The energy required to charge a parallel plate condenser of plate separation *d* and plate area of cross-section *A* such that the uniform electric field between the plates is E_1, is

[4 Marks]

(A) $(1 / 2) \varepsilon_0 E^2 Ad$

(B) $\varepsilon_0 E^2 / Ad$

(C) $\varepsilon_0 E^2 Ad$

(D) $(1 / 2) \varepsilon_0 E^2 / Ad$

Explanation:

Energy given by the cell $E = CV^2$

Here, C = capacitance of condenser = $(A\varepsilon_0 / d)$

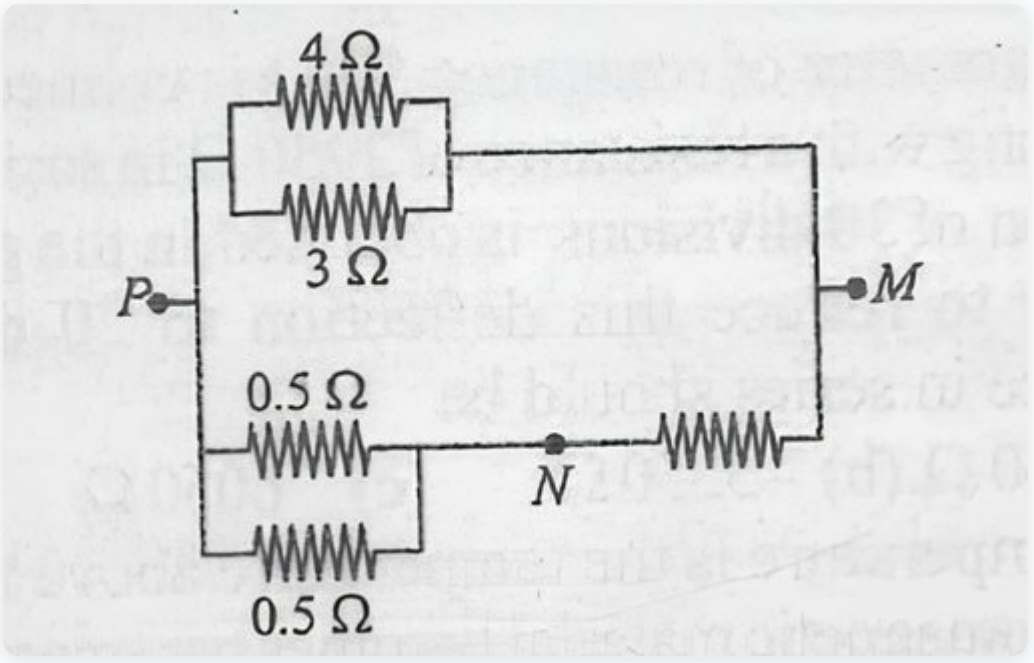
V = potential difference across the plates = Ed

$\therefore E = (A\varepsilon_0 / d) (Ed)^2 = A \varepsilon_0 E^2 d$

Question 91.

In the circuit shown, the current through the 4 Ω resistor is 1 A when the points P and M are connected to a DC voltage source. The potential difference between the points M and N is

[4 Marks]



(A) 1.0 V

(B) 1.5 V

(C) 0.5 V

(D) 3.2 V

Explanation:

Potential difference across PM

V = 4 × 1 = 4 volt (i.e., across 4 Ω)

Equivalent resistance of lower side arm,

R_eq = 1 + ((0.5 × 0.5) / (0.5 + 0.5)) = 1.25 Ω

Now the circuit can be shown as

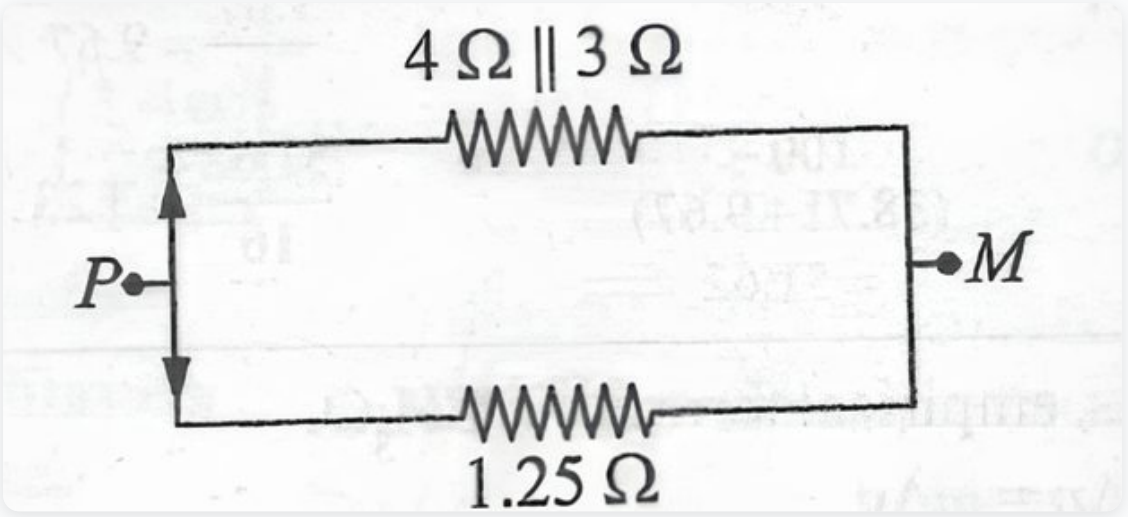
Let current I flows in lower branch, so

1.25 I = 4 V

or I = 4 / 1.25 = 3.2 A

Therefore, 3.2 A current flows in 1 Ω resistance, hence potential difference between M and N

V′ = 3.2 × 1 = 3.2 volt



Question 92.

What volume of oxygen gas (O_2) measured at 0°C and 1 atm, is needed to burn completely 1 L of propane gas (C_3H_8) measured under the same conditions?

[4 Marks]

(A) 5L

(B) 10 L

(C) 7L

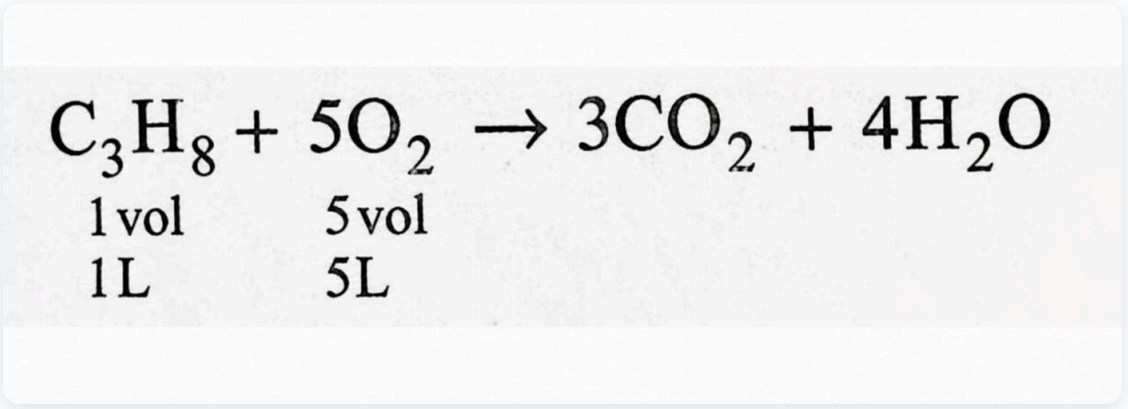
(D) 6L

Explanation:

Writing the equation of combustion of propane (C_3H_8), we get

From the above equation we find that we need 5 L of oxygen at N.T.P. to completely burn 1 L of propane at N.T.P.

If we change the conditions for both the gases from N.T.P. to same conditions of temperature and pressure, the same results are obtained. i.e. 5 L is the correct answer.



Question 93.

Number of moles of MnO_4^- required to oxidize one mole of ferrous oxalate completely in acidic medium will be:

[4 Marks]

(A) 7.5 moles

(B) 0.2 moles

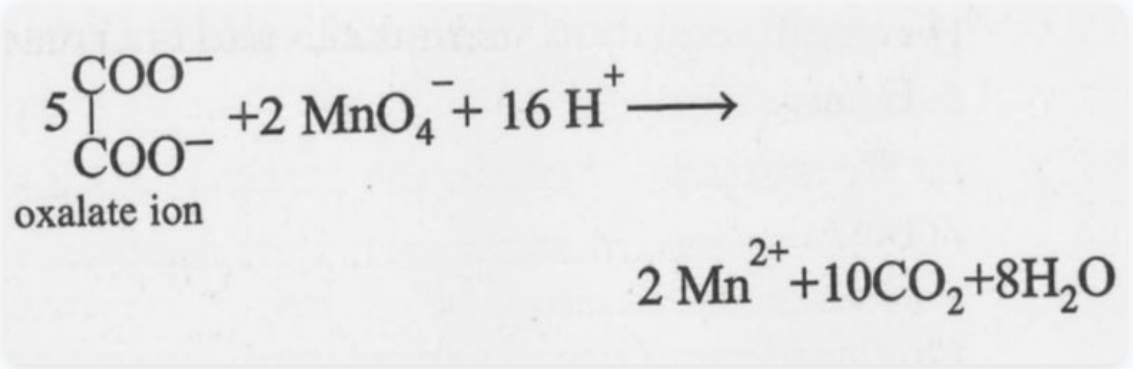
(C) 0.6 mole

(D) 0.4 moles

Explanation:

From above equation, 2 moles of MnO_4 required to oxidise 5 moles of oxalate.

Thus, number of moles of MnO_4^- required to oxidise one mole of oxalate = (2 × 2 / 5) = 0.4 moles.



Question 94.

Volume occupied by one molecule of water (density = 1 g cm^{-3}) is:

[4 Marks]

(A) 9.0×10^{-23} cm^3

(B) 3.0×10^{-23} cm^3

(C) 6.023×10^{-23} cm^3

(D) 5.5×10^{-23} cm^3

Explanation:

∴ Volume occupied by 1 gram water = 1 cm^3

or Volume occupied by

$(6.023 \times 10^{23} / 18)$ molecules of water = 1 cm^3

[∴ 1 g water = (1 / 18) moles of water]

Thus, volume occupied by 1 molecule of water

$= (1 \times 18 / 6.023 \times 10^{23}) cm^3 = 3.0 \times 10^{-23} cm^3$

i.e. the correct answer is option $(3.0 \times 10^{-23} cm^3)$.

Question 95.

An organic compound contains carbon, hydrogen and oxygen. Its elemental analysis gave C, 38.71% and H, 9.67%. The empirical formula of the compound would be:

[4 Marks]

(A) CH_2O

(B) CH_3O

(C) CHO

(D) CH_4O

Answer:

Element	%	Atomic weight	Atomic ratio	Simple ratio
C	38.71	12	$\frac{38.71}{12} = 3.23$	$\frac{3.23}{3.23} = 1$
H	9.67	1	$\frac{9.67}{1} = 9.67$	$\frac{9.67}{3.23} = 3$
O	$100 - (38.71 + 9.67) = 51.62$	16	$\frac{51.62}{16} = 3.23$	$\frac{3.23}{3.23} = 1$

Thus, empirical formula is CH_3O.

Question 96.

The measurement of the electron position if associated with an uncertainty in momentum, which is equal to 1×10^{-18} g cm s^{-1}. The uncertainty in electron velocity is, (mass of an electron is 9×10^{-28} g)

[4 Marks]

(A) 1×10^{11} cm s^{-1}

(B) 1×10^{9} cm s^{-1}

(C) 1×10^{5} cm s^{-1}

(D) 1×10^{6} cm s^{-1}

Explanation:

$\Delta p = m\Delta v$

Substituting the given values of Δx and m, we get

1×10^{-18} $gcms^{-1}$

$= 9 \times 10^{-28}$ g $\times \Delta v$

or $\Delta v = (1 \times 10^{-18} / 9 \times 10^{-28})$

$= 1.1 \times 10^{9}$ cm $s^{-1} \simeq 1 \times 10^{9}$ cm s^{-1}

Question 97.

How many moles of lead (II) chloride will be formed from a reaction between 6.5 g of PbO and 3.2 g of HCl?

[4 Marks]

(A) 0.029

(B) 0.011

(C) 0.044

(D) 0.333

Explanation:

Writing the equation for the reaction, we get

From this equation we find 223 g of PbO reacts 0with 73 g of HCl to form 278 g of $PbCl_2$.

If we carry out the reaction between 3.2 g HCl and 6.5 g PbO.

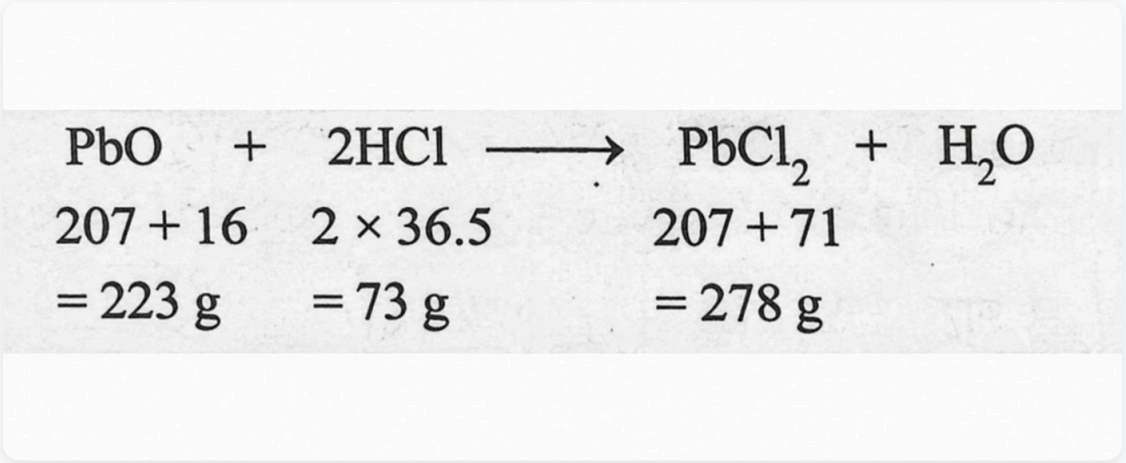
Amount of PbO that reacts with 3.2 g HCl = (223 / 73) × 3.2 = 9.77 g.

Since amount of PbO present is only 6.5 g so PbO is the limiting reagent.

Amount of $PbCl_2$ formed by 6.5 g of PbO = (278 / 223) × 6.5 g

Number of moles of $PbCl_2$ formed

= (278 / 223) × (6.5 / 278) moles = 0.029 moles.



Question 98.

If uncertainty in position and momentum are equal, then uncertainty in velocity is:

[4 Marks]

(A) √(h / (2π))

(B) (1 / 2m) √(h / π)

(C) (1 / m) √(h / π)

(D) √(h / π)

Explanation:

We know Δp·Δx ≥ h / 4π

or m·Δv·Δx = h / 4π

[∵ Δp = mΔv]

Since Δp = Δx (given)

∴ Δp·Δp = h / 4π or mΔv mΔv = h / 4π

or $(\Delta v)^2 = h / 4\pi m^2$

∴ $\Delta v = \sqrt{(h / 4\pi m^2)} = 1 / 2m \sqrt{(h / \pi)}$

Question 99.

The angular shape of ozone molecule (O_3) consists of:

[4 Marks]

(A) 2 sigma and 2 pi bonds

(B) 1 sigma and 1 pi bonds

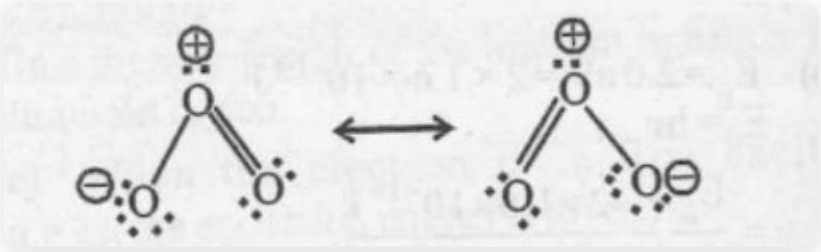
(C) 2 sigma and 1 pi bonds

(D) 1 sigma and 2 pi bonds

Explanation:

The shape of ozone molecules is

In it we find 2σ and 1π bond, i.e., the correct answer is 2 sigma and 1 pi bond



Question 100.

The correct order of increasing bond angles in the following triatomic species is:

[4 Marks]

(A) $NO_2^+ < NO_2 < NO_2^-$

(B) $NO_2^- < NO_2^+ < NO_2$

(C) $NO_2^- < NO_2 < NO_2^+$

(D) $NO_2^+ < NO_2^- < NO$

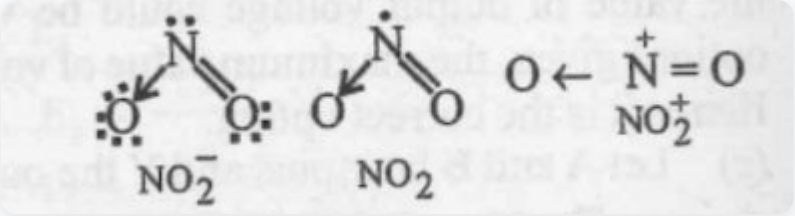
Explanation:

From the structure of three species we can determine the number of lone pair electrons on the central atom (i.e. N atom) and thus the bond angle.

We know that higher the number of lone pairs of electron on the central atom, greater is the lp–lp repulsion between Nitrogen and Oxygen atoms. Thus smaller is the bond angle.

The correct order of bond angle is

$NO_2^- < NO_2 < NO_2^+$



Question 101.

Which of the following are not the state functions?

(I) q + w

(II) q

(III) w

(IV) H − TS

[4 Marks]

(A) (I), (II) and (III)

(B) (II) and (III)

(C) (I) and (IV)

(D) (II), (III) and (IV)

Explanation:

We know that q (heat) and work (w) are not state functions but (q + w) is a state functions. H – TS (i.e. G) is also a state function.

Question 102.

Cornea transplant in humans is almost never rejected. This is because

[4 Marks]

(A) it is composed of enucleated cells

(B) it is a non-living layer

(C) it has no blood supply

(D) its cells are least penetrable by bacteria

Explanation:

Cornea transplant in humans is almost never rejected. This is because it has no blood supply. Cornea is a transparent layer of tissue, continuous with the sclerotic, that forms the front part of the vertebrate eye, over the iris and lens.

Question 103.

The blood calcium level is lowered by the deficiency of

[4 Marks]

(A) parathormone

(B) thyroxine

(C) both calcitonin and parathormone

(D) calcitonin

Explanation:

A peptide hormone secreted by the parathyroid gland in response to low levels of calcium in the blood. It acts to maintain normal blood levels of calcium by increasing the number of osteoclasts, which break down the bone matrix and release calcium into the blood. It also increases the reabsorption of calcium and magnesium ions in the kidney tubules, so that their concentration is maintained in the blood.

Question 104.

Which one of the following pairs of organs includes only the endocrine glands?

[4 Marks]

(A) Parathyroid and Adrenal

(B) Pancreas and Parathyroid

(C) Thymus and Testes

(D) Adrenal and Ovary

Explanation:

Parathyroid and adrenal are the endocrine glands. Adrenal glands are the pairs of endocrine glands situated immediately above the kidneys. Hence they are also known as suprarenal glands.

Parathyroid glands are the two pairs of endocrine glands situated behind or embedded within the thyroid gland in higher vertebrates. They produce parathyroid hormone, which controls the amount of calcium in the blood.

Question 105.

Unisexuality of flowers prevents

[3 Marks]

(A) both geitonogamy and xenogamy

(B) autogamy and geitonogamy

(C) geitonogamy, but not xenogamy

(D) autogamy, but not geitonogamy

Explanation:

Unisexuality of flowers prevents autogamy, but not geitonogamy. In self fertilisation, the male and female gametes are derived from the same individual. Among plants, self fertilization also called autogamy is common in many cultivated species, e.g., wheat and oats.

However, self fertilization is a form of inbreeding and does not allow for the mixing of genetic material; if it occurs over a number of generations it will result in offspring being less vigorous and productive than those resulting from cross fertilization.

Question 106.

What does the filiform apparatus do at the entrance into ovule?

[4 Marks]

(A) It helps in the entry of pollen tube into a synergid

(B) It guides pollen tube from a synergid to egg

(C) It brings about opening of the pollen tube

(D) It prevents entry of more than one pollen tube into the embryo sac

Explanation:

Filiform apparatus helps in the entry of pollen tube into a synergid in ovule.

Filiform apparatus is form of finger like projection comprising a core of micro fibrils enclosed in a sheath. The filiform apparatus resembles transfer cells meant for short distance movement of metabolites. The filiform apparatus is responsible for the absorption of food from the nucleus.

Question 107.

Endosperm is consumed by developing embryo in the seed of

[4 Marks]

(A) maize

(B) pea

(C) castor

(D) coconut

Explanation:

In pea, endosperm is consumed by developing embryo in the seed. The endosperm is completely absorbed by the growing embryo and the food reserve gets stored in the cotyledons. Such seeds are called non endospermic or exalbuminous.

Question 108.

Dry indehiscent single-seeded fruit formed from bicarpellary syncarpous inferior ovary is

[4 Marks]

(A) Berry

(B) Cremocarp

(C) Caryopsis

(D) Cypsela

Explanation:

Dry indehiscent single-seeded fruit formed from bicarpellary syncarpous inferior ovary is cypsela. Cypsela is also calied inferior, false or pseudocarpic achene, the thin fruit wall (developed from pericarp and thalamus)is attached to the seed at one point but the fruits develops from an inferior, unilocular and uniovuled ovary, e.g., sunflower, marigold. Some cypsela develop pappus for dispersal e.g. Soncus, Taraxacum.

Question 109.

The fleshy receptacle of syconus of fig encloses a number of:

[4 Marks]

(A) Mericarps

(B) Berries

(C) Samaras

(D) Achenes

Explanation:

The fleshy receptacle of syconus of fig encloses a number of Achenes. Achene is the thin dry pericarp is free from seed except at one point. The fruit develops from a monocarpellary pistil having superior unilocular and ovuled ovary.

Question 110.

The fruit is chambered, developed from inferior ovary and has seeds with succulent testa in

[4 Marks]

(A) guava

(B) cucumber

(C) orange

(D) pomegranate

Explanation:

The fruit of pomegranate is chambered, developed from inferior ovary and has seeds with succulent testa. The fruit is Balausta. It is a special type of false or pseudocarpic

berry in which the syncarpous pistil consists of two rows of fused carpels, one above the other.

Question 111.

Which one of the following statements is incorrect about menstruation?

[4 Marks]

(A) The beginning of the cycle of menstruation is called menarche

(B) At menopause in the female, there is especially abrupt increase in gonadotropic hormones

(C) During normal menstruation about 40 ml blood is lost

(D) The menstrual fluid can easily clot

Explanation:

In human female, the periodic discharge of blood, mucus and cellular debris from uterine wall from non pregnant women of sexual maturity is known as menstrual

cycle. Few hours before the start of menstruation, the spiral arterioles constrict one by one resulting into the blanched mucosa. The mucosa shrinks and the death of the blood deprived tissues takes place. Blood clotting doesn't occur due to presence of fibrinolysin.

Question 112.

Which extra-embryonic membrane in humans prevents desiccation of the embryo inside the uterus?

[4 Marks]

(A) Chorion

(B) Allantois

(C) Yolk sac

(D) Amnion

Explanation:

An organizer is a signal aligning center that directs the development of embryo or part of embryo, e.g. in amphibians, it is dorsal tip of blastophore.

Question 113.

Given below are four methods (A-D) and their modes of action (a-d) in achieving contraception. Select their correct matching from the four options that follow:

[4 Marks]

Method	Mode of Action
A. The pill	(a) Prevents sperms reaching cervix
B. Condom	(b) Prevents implantation
C. Vasectomy	(c) Prevents ovulation
D. Copper T	(d) Semen contains no sperms

(A) A -(b), B-(c), C-(a), D-(d)

(B) A -(d), B-(a), C-(b), D-(c)

(C) A -(c), B-(d), C-(a), D-(b)

(D) A -(c), B-(a), C-(d), D-(b)

Explanation:

A. The pill — Prevents ovulation

B. Condom — Prevents sperm reaching cervix

C. Vasectomy — Semen contains nosperms

D. Copper-T — Prevent implantation

Question 114.

Consider the statements given below regarding contraception and answer as directed thereafter

(i) Medical Termination of Pregnancy (MTP) during first trimester is generally safe

(ii) Generally chances of conception are nil until mother 'breast-feeds the infant upto two years

(iii) Intrauterine devices like copper-T are effective contraceptives

(iv) Contraception pills may be taken upto one week after coitus to prevent conception

Which two of the above statements are correct?

[4 Marks]

(A) iii and iv

(B) ii and iii

(C) i and ii

(D) i and iii

Explanation:

Statements i and iii are correct.

- Medical Termination of Pregnancy (MTP) during first trimester is generally safe.
- Intrauterine device like copper-T are effective contraceptives.

Question 115.

Which one of the following conditions in humans is correctly matched with its chromosomal abnormality/linkage?

[4 Marks]

(A) Downs syndrome – 44 autosomes + XO

(B) Erythroblastosis foetalis – X-linked

(C) Klinefelter's syndrome – 44 autosomes + XXY

(D) Colour blindness – Y-linked

Explanation:

Klinefelter's syndrome is a genetic disorder affecting men in which an individual gains an extra X chromosome, so that the usual karyotype of XY is replaced by XXY. Symptoms of Klinefelter, syndrome named after U.S. physician H.P. Klinefelter, include female characteristics (such as breast enlargement).

Question 116.

Haploids are more suitable for mutation studies than the diploids. This is because

[4 Marks]

(A) haploids are reproductively more stable than diploids

(B) mutagens penetrate in haploids more effectively than diploids

(C) haploids are more abundant in nature than diploids

(D) all mutations, whether dominant or recessive are expressed in haploids

Explanation:

Haploids describes a nucleus cell or organism with a single set of unpaired chromosomes. The haploid number is designated as X. Reproductive cells, formed as a result of meiosis are haploid. Fusion of two such cells restores the normal diploid number. Therefore, haploids are more suitable for mutation studies than the diploids. This is because all mutations, whether dominant or recessive are expressed in haploids.

Question 117.

Which one of the following pairs of nitrogenous bases of nucleic acids, is wrongly matched with the category mentioned against it?

[4 Marks]

(A) Uracil, Cytosine – Pyrimidines

(B) Thymine, Uracil – Pyrimidines

(C) Adenine, Thymine – Purines

(D) Guanine, Adenine – Purines

Explanation:

Purine is an organic nitrogenous base sparingly soluble in water, that gives rise to a group of biologically important derivatives, notably adenine and guanine, which occur in nucleotides and nucleic acids (DNA and RNA).

Question 118.

Which one of the following pairs of codons is correctly matched with their function or the signal for the particular amino acid?

[4 Marks]

(A) UUA, UCA – Leucine

(B) AUG, ACG – Start/Methionine

(C) UAG, UGA – Stop

(D) GUU, GCU – Alanine

Explanation:

GCU indicates alanine but GUU indicates valine. Stop codons are AUG, UGA and UAA.

AUG is the most common start codon which codes for methionine.

UUA indicates leucine but UCA indicates serine.

Question 119.

Polysome is formed by

[4 Marks]

(A) a ribosome with several subunits

(B) ribosomes attached to each other in a linear arrangement

(C) several ribosomes attached to a single mRNA

(D) many ribosomes attached to a strand of endoplasmic reticulum

Explanation:

Polysome (Polyribosome) is a complex formed by several ribosomes attached to a single mRNA molecule in the process of translation.

Question 120.

In the DNA molecule

[4 Marks]

(A) there are two strands which run parallel in the 5′ → 3′ direction

(B) the proportion of Adenine in relation to thymine varies with the organism

(C) there are two strands which run anti-parallel, one in 5′ → 3′ direction and other in 3′ → 5′

(D) the total amount of purine nucleotides and pyrimidine nucleotides is not always equal

Explanation:

In the DNA molecule, there are two strands which run anti-parallel, one in 5'–3' direction and the other in 3'–5' direction. The two strands are held together by hydrogen bonds between their bases. Adenine (A), a purine of one chain is exactly opposite to thymine (T), a pyrimidine of the other chain. Similarly, cytosine (C), a pyrimidine lies opposite to guanine (G), a purine. This allows a sort of lock & key arrangement between large sized purine & small sized pyrimidine. It is strengthened by the appearance of hydrogen bonds between the two.

Question 121.

Which one of the following scientists name is correctly matched with the theory put forth by him?

[4 Marks]

(A) Pasteur – Inheritance of acquired characters

(B) Weismann – Theory of continuity of Germplasm

(C) De Vries – Natural selection

(D) Mendel – Theory of Pangenesis

Explanation:

The theory of the continuity of the germplasm published by August Weismann (1834–1914) in 1886. It proposes that the contents of the reproductive cells (sperms and ova) are passed on unchanged from

one generation to the next, unaffected by any changes undergone by the rest of the body. It thus rules out any possibility of the inheritance of acquired characteristics, and has become fundamental to Neo-Darwinian theory.

Question 122.

Darwinism finches are an excellent example of

[4 Marks]

(A) connecting links

(B) brood parasitism

(C) seasonal migration

(D) adaptive radiation

Explanation:

Darwin's Finches are an excellent example of adaptive radiation. Adaptive radiation is the occurrence of the diversification of a single or small groups of species into a large

number or descendents that occupy various ecological niches. This is an evolutionary process driven by natural selection.

Question 123.

Thorn of Bougainvillea and tendril of Cucurbita are example of

[4 Marks]

(A) analogous organs

(B) homologous organs

(C) vestigial organs

(D) retrogressive evolution

Explanation:

Thorn of Bougainvillea and tendril of Cucurbita are example of homologous organs.

Question 124.

Which one of the following pairs of items correctly belongs to the category of organs mentioned against it?

[4 Marks]

(A) Wings of honey bee and wings of crow – Homologous organs

(B) Nephridia of earthworm and malpighian tubules of Cockroach – Excretory organs

(C) Nictitating membrane and blind spot in human eye – Vestigial organs

(D) Thorn of Bougainvillea and tendril of Cucurbita – Analogous organs

Explanation:

Nephridia of earthworm and malphigian tubules of cockroach belongs to excretory organs. Earthworm has a well-developed excretory system which is composed of large number of minute, coiled and glandular segmentally arranged excretory tubules called the nephridia. Malphigian tubules of cockroach are extremely fine yellowish unbranched thread-like structure present at the function of midgut and hindgut.

Question 125.

Which one of the following is incorrect about the characteristics of protobionts (coacervates and microspheres) as envisaged in the biogenic origin of life?

[4 Marks]

(A) They were able to reproduce

(B) They could separate combinations of molecules from the surroundings

(C) They could maintain an internal environment

(D) They were partially isolated from the surroundings

Explanation:

Coacervate is an aggregate of macromolecules such as proteins, lipids and membrane, and contain enzymes that are capable of converting substances such as glucose into more complex molecules, such as starch. Coacervate droplets arise spontaneously under appropriate conditions and may have been the prebiological systems from which living organisms originated.

Question 126.

Match the disease in Column I with the appropriate items (pathogen/prevention/treatment) in Column II.

[4 Marks]

Column I	Column II
1. Amoebiasis	(i) Treponema pallidum
2. Diphtheria	(ii) Use only sterilized food and water
3. Cholera	(iii) DPT Vaccine
4. Syphilis	(iv) Use oral rehydration therapy

(A) (1) – (i), (2) – (ii), (3) – (iii), (4) – (iv)

(B) (1) – (ii), (2) – (iv), (3) – (i), (4) – (iii)

(C) (1) – (ii), (2) – (i), (3) – (iii), (4) – (iv)

(D) (1) – (ii), (2) – (iii), (3) – (iv), (4) – (i)

Explanation:

Amoebiasis : Use only sterilized food and water

Diphtheria : DPT Vaccine

Cholera : Use oral rehydration therapy

Syphilis : Treponema pallidum

Question 127.

To which type of barriers under innate immunity, do the saliva in the mouth and the tears from the eyes, belong?

[4 Marks]

(A) Cytokine barriers

(B) Cellular barriers

(C) Physiological barriers

(D) Physical barriers

Explanation:

The saliva in the mouth and the tears from the eyes belong to physiological barriers.

Question 128.

Which one of the following is the correct statement regarding the particular psychotropic drug specified?

[4 Marks]

(A) Opium stimulates nervous system and causes hallucinations

(B) Hashish causes after thought perceptions and hallucinations

(C) Morphine leads to delusions and disturbed emotions

(D) Barbiturates cause relaxation and temporary euphoria

Explanation:

Hashish causes after-thought perceptions and hallucinations. Hashish is a preparation of Cannabis composed of the compressed trichomes collected from the Cannabis plant. Psychoactive effects vary between types of Hashish but are usually the same as those of other Cannabis preparations such as marijuana. Hashish is generally prohibited to the same extent as all other forms of cannabis. It is consumed in much the same way as Cannabis buds, used by itself in miniature smoking pipes, vapourized, hot-knifed, or smoked in joints mixed with tobacco, Cannabis buds or other herbs.

Question 129.

Which one of the following is being tried in India as a biofuel substitute for fossil fuels?

[4 Marks]

(A) Aegilops

(B) Musa

(C) Azadirachta

(D) Jatropha

Explanation:

Jatropha is being tried in India as a biofuel substitute for fossil fuels. Biofuels are the fuels of biological origin. They consist of renewable form of energy.

Energy problems can be solved up to a great extent by efficient and proper used of these biofuels.

Question 130.

Consider the following four measures (a–d) that could be taken to successfully grow chickpea in an area where bacterial blight disease is common

(1) Spray with Bordeaux mixture

(2) Control of the insect vector of the disease pathogen

(3) Use of only disease-free seeds

(4) Use of varieties resistant to the disease

Which two of the above measures can control the disease?

[4 Marks]

(A) 1 and 4

(B) 2 and 3

(C) 1 and 2

(D) 3 and 4

Explanation:

Use of only disease free seeds and use of disease resistant varieties are the most important control measures that could be taken to successfully grow chickpea in an

area where bacterial blight disease is common.

Question 131.

Which one of the following proved effective for biological control of nematodal disease in plants?

[4 Marks]

(A) Gliocladium virens

(B) Pseudomonas cepacia

(C) Pisolithus tinctorius

(D) Paecilomyces lilacinus

Explanation:

Paecilomyces is a cosmopolitan filamentous fungus which inhabits the soil, decaying plants, and food products. Some species of Paecilomyces are isolated from insect. The telomorphs of Paecilomyces are classified in the genera Cilromocleista, Talaromyces. Paecilomyces is usually considered as a contaminant but may also cause

infections in humans and animals.

Question 132.

Trichoderma harzianum has proved a useful microorganism for

[4 Marks]

(A) reclamation of wastelands

(B) gene transfer in higher plants

(C) biological control of soil-borne plant pathogens

(D) bioremediation of contaminated soils

Explanation:

Trichoderma harzianum has proved a useful microorganism for biological control of soil borne plant pathogens. Trichoderma harzianum is a fungus that is also used as a fungicide. It is used for foliar application, seed treatment and soil treatment for suppression of various disease causing fungal pathogens. Commercial biotechnological products such as 3 Tac have been useful for treatment of Botrytis, Fusarium, Penicillium sp. It is also used for manufacturing enzymes.

Question 133.

Modern detergents contain enzyme preparations of:

[4 Marks]

(A) Acidophiles

(B) Alkaliphiles

(C) Thermoacidophiles

(D) Thermophiles

Explanation:

Modern detergent contain enzymes preparations of alkaliphiles. Alkaliphiles live in alkaline environments like soda lakes, or alkaline soil. The pH level in alkaline substances is from about 9 to 11 on the pH scale. The alkaliphiles keep an alkaline level of about pH 9 inside the cells, but the environment has a higher level of alkalinity. Alkaliphiles can keep the right pH level by pumping in hydrogen ions in out at the right rate.

Question 134.

Human insulin is being commercially produced from a transgenic species of

[4 Marks]

(A) Escherichia

(B) Mycobacterium

(C) Saccharomyces

(D) Rhizobium

Explanation:

Human insulin is being commercially produced from a transgenic species of *Escherichia coli*. *E. coli* is a bacterium that is commonly found in the lower intestine of warm blooded animals. The bacteria can also be grown easily and its genetics are comparatively simple and easily manipulated, making it one of the best studied prokaryotic model organisms, and an important species in biotechnology.

Question 135.

Which one of the following is linked to the discovery of Bordeaux mixture as a popular fungicide?

[4 Marks]

(A) Black rust of wheat

(B) Loose smut of wheat

(C) Downy mildew of grapes

(D) Bacterial leaf blight of rice

Explanation:

Downy mildew of grapes disease is linked to the discovery of Bordeaux mixture as a popular fungicide. Bordeaux mixture was discovered by Millardet in France in 1882. Bordeaux mixture is prepared by dissolving 40g of copper sulphate and 40 g of calcium hydroxide in 5 litres of water. Bordeaux mixture is used primarily as a fungicide, it was first used to control downy mildew disease of grape-vine caused by a fungus, Plasmopara viticola.

Question 136.

The linking of antibiotic resistance gene with the plasmid vector became possible with

[4 Marks]

(A) DNA polymerase

(B) Endonucleases

(C) DNA ligase

(D) Exonucleases

Explanation:

The linking of antibiotic resistance gene with the plasmid vector became possible with DNA ligase. DNA ligase is an enzyme that is able to join together two portions of DNA and therefore plays an important role in DNA repair. DNA ligase is also used in recombinant DNA technology as it ensures that the foreign DNA is bound to the plasmid into which it is incorporated.

Question 137.

Gel electrophoresis is used for

[4 Marks]

(A) isolation of DNA molecule

(B) construction of recombinant DNA by joining with cloning vectors

(C) cutting of DNA into fragments

(D) separation of DNA fragments according to their size

Explanation:

Gel electrophoresis is a technique far separation of DNA fragments according to their size. DNA is negatively charged so in gel tank when electricity is passed, DNA move towards positive electrode.

Question 138.

What is antisense technology?

[4 Marks]

(A) RNA polymerase producing DNA

(B) When a piece of RNA that is complementary in sequence is used to stop expression of a specific gene

(C) Production of somaclonal variants in tissue cultures

(D) A cell displaying a foreign antigen used for synthesis of antigens

Explanation:

An RNA molecule whose base sequence is complementary to that of the RNA transcript of a gene, i.e. the sense RNA, such as a messenger RNA (mRNA). Hence, an antisense RNA can undergo base pairing with its complementary mRNA sequence. This blocks gene expression, either by preventing

access for ribosome to translate the mRNA or by triggering degradation of the double stranded RNA by ribonuclease

enzymes. Antisense RNA and DNA both have therapeutic potential for modifying the activity of disease causing genes.

Question 139.

Cry endotoxins obtained from Bacillus thuringiensis are effective against

[4 Marks]

(A) nematodes

(B) boll worms

(C) flies

(D) mosquitoes

Explanation:

Cry endotoxin obtained from Bacillus thuringiensis are effective against bollworms. A bollworm is a common term for any larva of a moth that attacks the fruiting bodies

of certain crops, especially cotton.

Question 140.

Consider the following statements about biomedical technologies?

1. During open heart surgery blood is circulated in the heart-lung machine
2. Blockage in coronary arteries is removed by angiography
3. Computerised axial tomography (CAT) shows detailed internal structure as seen in a section of body
4. X-ray provides clear and detailed images of organs like prostate glands and lungs Which two of the above statements are correct?

[4 Marks]

(A) 3 and 4

(B) 2 and 4

(C) 1 and 2

(D) 1 and 3

Explanation:

A modern hospital can make use of variety of sophisticated instruments and equipment for accurate diagnosis and treatment of diseases. Three main categories of instruments and equipments used are diagnostic, imaging and therapeutic. From the above given statements, statement 1 & 3 are correct.

- During open heart surgery blood is circulated in the heart lung machine.
- Computerised Axial Tomography (CAT) shows detailed internal structure as seen in a section of body.

Question 141.

A transgenic food crop which may help in solving the problem of night blindness in developing countries is

[4 Marks]

(A) Bt Soybean

(B) Golden rice

(C) Starlink maize

(D) Flavr Savr tomatoes

Explanation:

Golden rice is a transgenic food crop which may help in solving the problem of night blindness in developing countries. Golden rice or miracle rice is rich in vitamin A or β-carotene and iron.

Question 142.

Consider the following four statements (a–d) about certain desert animals such as kangaroo rat.

1. They have dark colour and high rate of reproduction and excrete solid urine
2. They do not drink water, breathe at a slow rate to conserve water and have their body covered with thick hairs
3. They feed on dry seeds and do not require drinking water
4. They excrete very concentrated urine and do not use water to regulate body temperature

Which two of the above statements for such animals are true?

[4 Marks]

(A) 2 and 3

(B) 3 and 4

(C) 1 and 2

(D) 3 and 1

Explanation:

Characteristics of certain desert animals such as kangaroo, rat are

- They feed on dry seeds and do not require drinking water.
- They excrete very concentrated urine and do not use water to regulate body temperature.

Question 143.

What is true about the isolated small tribal populations?

[4 Marks]

(A) There is a decline in population as boys marry girls only from their own tribe.

(B) Hereditary diseases like colour blindness do not spread in the isolated population.

(C) Wrestlers who develop strong body muscles in their lifetime pass this character on to their progeny.

(D) There is no change in population size as they have a large gene pool.

Explanation:

'A' is more recent and shows slight reduction in growth rate.

Question 144.

About 70% of total global carbon is found in

[4 Marks]

(A) agroecosystems

(B) grasslands

(C) oceans

(D) forests

Explanation:

About 70% of total global carbon is found in oceans. This oceanic reservoir regulates the amount of carbon dioxide in the atmosphere. Atmosphere contains only about one percent of total global carbon.

Question 145.

The slow rate of decomposition of fallen logs in nature is due to their

[4 Marks]

(A) low moisture content

(B) anaerobic environment around them

(C) low cellulose content

(D) poor nitrogen content

Explanation:

Low moisture content will lead to decrease in number of decomposers because decomposers need an optimum moisture for their growth and functioning.

Question 146.

Consider the following statements concerning food chains

1. removal of 80% tigers from an area resulted in greatly increased growth of vegetation
2. removal of most of the carnivores resulted in an increased population of deers
3. the length of food chains is generally limited to 3–4 trophic levels due to energy loss
4. the length of food chains may vary from 2 to 8 trophic levels

Which two of the above statements are correct?

[4 Marks]

(A) (3) and (4)

(B) (1) and (4)

(C) (2) and (3)

(D) (1) and (2)

Explanation:

Food chain is the transfer of energy from green plants (Primary producers), through a sequence of organisms occupies in a food chain is known as its trophic level. Therefore, statements b and c are

correct.

Question 147.

The table below gives the population (in thousands) of ten species (A–J) in four areas (a–d) consisting of the number of habitats given within brackets against each. Study the table and answer the question which follows:

Which area out of a to d shows maximum species diversity?

[4 Marks]

Area and Number of habitats	Species and their populations (in thousands) in the									
	A	B	C	D	E	F	G	H	I	J
a (11)	2.3	1.2	0.52	6.0	_	3.1	1.1	9.0	_	10.3
b (11)	10.2	_	0.62	_	1.5	3.0	_	8.2	1.1	11.2
c (13)	11.3	0.9	0.48	2.4	1.4	42	0.8	8.4	2.2	4.1
d (12)	3.2	10.2	11.1	4.8	0.4	33	0.8	7.3	113	2.1

(A) a

(B) d

(C) c

(D) b

Explanation:

Species richness refers to the number of species per unit area. Species evenness is the relative abundance with which each species is represented in an area. Thus variation in the number of species, kinds of species as well as the number of individuals per species lead to diversity. Area(d) and number of habitats (12) shows the maximum species diversity.

Question 148.

Quercus species are the dominant component in

[4 Marks]

(A) Alpine forests

(B) Tropical rain forests

(C) Scrub forests

(D) Temperate deciduous forests

Explanation:

Quercus species is a dominant component in temperate - deciduous forests. The common name of Quercus is cork oak belongs to family fabaceae. Cork used as a bottle stoppers, soles for shoes, insulating material and for manufacture of linoleum.

Question 149.

Which one of the following is not observed in biodiversity hotspots?

[4 Marks]

(A) Accelerated species loss

(B) Species richness

(C) Lesser inter-specific competition

(D) Endemism

Explanation:

Lesser inter-specific competition is not observe in biodiversity hotspots, levels of species richness and high degree of endemism. Initially 25 biodiversity hotspots

were identified but subsequently nine more have of biodiversity hotspots in the world to 34. These hotspots are also regions of accelerated habitat loss.

Question 150.

World Summit on Sustainable Development (2002) was held in

[4 Marks]

(A) South Africa

(B) Argentina

(C) Sweden

(D) Brazil

Explanation:

World Summit on sustainable development (2002) was held in Johannesburg, South Africa, 190 countries pledged their commitment to achieve by 2010, a significant in reduction the current rate of biodiversity loss at global, regional and local levels.

Question 151.

According to Central Pollution Control Board (CPCB), which particulate size in diameter (in micrometers) of the air pollutants is responsible for greatest harm to human health?

[4 Marks]

(A) 1.0 or less

(B) 5.2 – 2.5

(C) 2.5 or less

(D) 1.5 or less

Explanation:

The Central Pollution Control Board (CPCB), statutory organisation, was constituted in September 1974 under the Water (Prevention and Control of Pollution) Act 1974. Further, CPCB was entrusted with the powers and functions under Air (Prevention and Control of Pollution) Act, 1981. It serves as a field formation and also provides technical services to the Ministry of Environment and Forest, of the provisions of the Environment (Protection) Act, 1986. Principal functions of the CPCB, as spelt out in the Water (Prevention and Control of Pollution) Act, 1974, and the Air (Prevention and Control of Pollution) Act, 1981.

(i) To promote cleanliness of streams and wells in different areas of the states by prevention, control and abatement of water pollution.

(ii) To improve the quality of air and to prevent, control or abate air pollution in the country.

Question 152.

A lake near a village suffered heavy mortality of fishes within a few days. Consider the following reasons for this?

1. Lots of urea and phosphate fertilizer were used in the crops in the vicinity
2. The area was sprayed with DDT by an aircraft
3. The lake water turned green and stinky
4. Phytoplankton populations in the lake declined initially thereby greatly reducing photosynthesis.

Which two of the above were the main causes of fish mortality in the lake ?

[4 Marks]

(A) (1) and (3)

(B) (3) and (4)

(C) (1) and (2)

(D) (1) and (4)

Explanation:

Lots of urea and phosphate fertilizer were used in the crops in the vicinity and the lake water turned green and stinky. Due to this, lake near a village suffered heavy mortality of fishes within a few days.

Question 153.

Which one of the following is the correct percentage of the two (out of the total of 4) green house gases that contribute to the total global warming?

[4 Marks]

(A) CFCs 14%, Methane 20%

(B) CO_2 40%, CFCs 30%

(C) N_2O 6%, CO_2 86%

(D) Methane 20%, N_2O 18%

Explanation:

CFC 14%, Methane 20% is the correct percentage of the two (out of the total 4) greenhouse gases that contribute to the total global warming. A regular assessment of abundance of greenhouse gases and their impact on global environment is being made by IPCC (Inter-Governmental Panel on Climate

Change). The various greenhouse gases are CO_2 (warming effect 60%), CH_4 (effect 20%), chlorofluorocarbons or CFCs (14%) and nitrous oxide (N_2O, 6%). Others of minor significance are water vapours and ozone.

Question 154.

In human adult females oxytocin:

[4 Marks]

(A) stimulates pituitary to secrete vasopressin

(B) is secreted by anterior pituitary

(C) causes strong uterine contractions during parturition

(D) stimulates growth of mammary glands

Explanation:

In human adult females, oxytocin is a hormone released by the pituitary gland (neurohypophysis), that causes contraction of uterine muscles during labour.

Question 155.

What is vital capacity of our lungs?

[4 Marks]

(A) Total lung capacity minus expiratory reserve volume

(B) Total lung capacity minus residual volume

(C) Inspiratory reserve volume plus expiratory reserve volume

(D) Inspiratory reserve volume plus tidal volume

Explanation:

Vital capacity of lungs is the largest possible expiration after largest possible inspiration that is greatest. Volume of air can be exchanged in single respiration or amount of air breathed in and out with greatest possible efforts.

Vital capacity = TV + IRV + ERV

Total lung capacity = VC + RV

So, we can say that VC = Total lung capacity - Residual volume

Question 156.

The C4 plants are photosynthetically more efficient than C3 plants because:

[4 Marks]

(A) CO_2 generated during photorespiration is trapped and recycled through PEP carboxylase

(B) the CO_2 compensation point is more

(C) the CO_2 efflux is not prevented

(D) they have more chloroplasts

Explanation:

C4plants are photosynthetically more efficient than C3 plants because C4 plant contain two types of chloroplast i.e., bundle sheath chloroplast and mesophyll chloroplast. So such plants operate a dicarboxylic acid cycle in addition to Calvin

cycle. CO_2, acceptor molecule (PEP) is present in large bundle sheath cell which has higher efficiency in picking up CO_2,.

Thus, photosynthesis continues even at low CO_2, concentration and the rate of photorespiration is also negligible.

Question 157.

Which one of the following is incorrect about the characteristics of protobionts (coacervates and microspheres) as envisaged in the abiogenicorigin of life?

[4 Marks]

(A) They could separate combinations of molecules from the surroundings

(B) They were able to reproduce

(C) They were partially isolated from the surruounding

(D) They could maintain an internal environment

Explanation:

Coacervates and microspheres reproduced by asexual methods such as budding, fission and fragmentation.

Question 158.

Bacterial leaf blight of rice is caused by aspecies of:

[4 Marks]

(A) Pseudomonas

(B) Xanthomonas

(C) Alternaria

(D) Erwinia

Explanation:

Bacterial leaf blight of rice is caused by Xanthomonasoryzae a bacterium which is gram- negative, aerobic, capsulated, and motile with a single polar flagellum. Primary

infection is carried through the infected seeds. The entry of the pathogen occurs through wounds and stomata. The symptoms of the disease are the appearance of linear, yellow to straw coloured stripes, usually on both the edges of the leaf. As the disease progresses, the drying and twisting of the leaf tip occurs. The most destructive phase of the disease is the 'Ikresek' or wilt resulting from early systematic infection.

Question 159.

If a gas expands at constant temperature, it indicates that:

[4 Marks]

(A) number of the molecules of gas increases

(B) kinetic energy of molecules remains the same

(C) pressure of the gas increases

(D) kinetic energy of molecules decreases

Explanation:

At any constant temperature the K.E. of gaseous molecules remains same.

Question 160.

Bond dissociation enthalpy of H_2, Cl_2 and HCl are 434, 242 and 431 kJ mol^{-1} respectively. Enthalpy of formation of HCl is:

[4 Marks]

(A) 245 kJ mol^{-1}

(B) –93 kJ mol^{-1}

(C) –245 kJ mol^{-1}

(D) 93 kJ mol^{-1}

Explanation:

The reaction for formation of HCl can be written as

$H_2 + Cl_2 \rightarrow 2HCl$

H–H + Cl–Cl → 2(H–Cl)

Substituting the given values, we get enthalpy of formation of

$\Delta H_f(HCl) = [(B.E)_{H–H} + (B.E)_{Cl–Cl}] – [2 \times B.E_{(H–Cl)}]$

= 434 + 242 – 2 × (431) = –186 kJ

∴ Enthalpy of formation for

1 mol HCl = (–186 / 2) kJ = –93 kJ

Question 161.

For the gas phase reaction,

$PCl_5(g) \rightleftharpoons PCl_3(g) + Cl_2(g)$

which of the following conditions are correct?

[4 Marks]

(A) $\Delta H < 0$ and $\Delta S < 0$

(B) $\Delta H = 0$ and $\Delta S < 0$

(C) $\Delta H > 0$ and $\Delta S < 0$

(D) $\Delta H > 0$ and $\Delta S > 0$

Explanation:

For the reaction:

$PCl_5(g) \rightleftharpoons PCl_3(g) + Cl_2(g)$

The reaction given is an example of decomposition reaction and we know that decomposition reactions are endothermic in nature, i.e. $\Delta H > 0$.

Further

$\Delta n = (1 + 1) - 1 = +1$

Hence more number of molecules are present in products which shows more randomness i.e. $\Delta S > 0$ (ΔS is positive).

Question 162.

If the concentration of OH^- ions in the reaction

$Fe(OH)_3(s) \rightleftharpoons Fe^{3+}(aq) + 3OH^-(aq)$

is decreased by 1/4 times, then equilibrium concentration of Fe^{3+} will increase by:

[4 Marks]

(A) 16 times

(B) 4 times

(C) 8 times

(D) 64 times

Explanation:

For this reaction K_eq is given by

$K = [Fe^{3+}][OH^-]^3 / [Fe(OH)_3]$

$= (Fe^{3+})(OH^-)^3$ [$\because$ solid = 1]

If (OH^-) is decreased by 1 / 4 times, then for given reaction, equilibrium constant to remain constant, we have to increase the concentration of (Fe^{3+}) by a factor of (4^3) i.e. $4 \times 4 \times 4 = 64$. Thus, option (64 times) is correct.

Question 163.

The dissociation equilibrium of a gas AB_2 can be represented as

$2AB_2(g) \rightleftharpoons 2AB(g) + B_2(g)$

The degree of dissociation is 'x' and is small compared to 1. The expression relating the degree of dissociation (x) with equilibrium constant K_p and total pressure P is:

[4 Marks]

(A) (K_p / P)

(B) $(2K_p / P)^{1/3}$

(C) $(2K_p / P)^{1/2}$

(D) $(2K_p / P)$

Explanation:

For the reaction

$$2\,AB_2(g) \rightleftharpoons 2\,AB\,(g) + B_2(g)$$

initially	2	0	0
at equi	2(1 -x)	2x	x

The partial pressure at equilibrium are calculated on the basis of total numbers of moles at equilibrium.

Total number of moles

$= 2\,(1- x) + 2x + x = (2 + x)$

$\therefore P_AB_2 = (2(1 - x) / (2 + x)) \times P$ where P is the total pressure.

$P_AB = (2x / (2 + x)) \times P$, $P_B_2 = (x / (2+x)) \times P$

Since x is very small so can be neglected in denominator Thus, we get

$P_AB_2 = (1 - x) \times P$ $\qquad$ $P_AB = x \times P$

$P_B_2 = x / 2 \times P$

Now, $K_P = ((P_AB)^2 (P_B_2) / (P_AB_2)^2) = ((x)^2 \times P^2 \cdot P \times x / 2 / (1-x)^2 \times P^2)$

$= (x^3 \cdot P^3 / 2 \times 1 \times P^2)$ $\qquad$ $[\because 1 - x \simeq 1]$

$= x^3 \cdot P / 2$ or $x^3 = (2 \cdot Kp / P)$ or $x = (2K_P / P)^{1/3}$

Question 164.

The values of Kp_1 and Kp_2 for the reactions:

$X \rightleftharpoons Y + Z$ ...(a)

$A \rightleftharpoons 2B$ ...(b)

are in the ratio of 9 : 1. If degree of dissociation of X and A be equal, then total pressure at equilibrium for (a) and (b) are in the ratio:

[4 Marks]

(A) 1 : 9

(B) 36 : 1

(C) 1 : 1

(D) 3 : 1

Explanation:

Given reaction are

$x \rightleftharpoons Y+Z$ ...(i)

and $A \rightleftharpoons 2B$ ...(ii)

Let the total pressure for reaction (i) and (ii) be P_1 and P_2 respectively, then

$Kp_1/Kp_2 = 9/1$ (given)

After dissociation

$x \rightleftharpoons Y+Z$

$(1-\alpha)$ α α

At equilibrium

[Let 1 mole of X dissociate with α as degree of dissociation]

Total number of moles = $1-\alpha + \alpha+ \alpha = (1+\alpha)$

Thus $P_x = (1-\alpha/1+\alpha)\cdot P_1$; P_y= $(\alpha/1+\alpha) P_1$;

P_z = $(\alpha/1+\alpha)\cdot P_1$

$\therefore Kp_1= (\alpha/1+\alpha)\cdot P_1 \times \alpha/(1+\alpha)\cdot$

$P_1/(1-\alpha/1+\alpha)\cdot P_1$ ....(i)

Similarly for $A \rightleftharpoons 2B$

At equilibrium (1-α) 2α

We have,

$Kp_2 = (2\alpha P_2/1+\alpha)^2 / (1-\alpha/1+\alpha)\ P_2$ ....(ii)

Diving (i) by (ii), we get

$Kp_1/Kp_2 = \alpha^2 \cdot P_1/4\alpha^2 \cdot P_2$ or $Kp_1/Kp_2 = 1/4 \cdot P_1/P_2$

or $9 = 1/4 \cdot P_1/P_2$ [$\therefore Kp_1/Kp_2 = 9/1$]

or $P_1/P_2 = 36/1$ or $P_1 : P_2 = 36:1$

Question 165.

Equal volumes of three acid solutions of pH3, 4, and 5 are mixed in a vessel. What will be the H^+ ion concentration in the mixture?

[4 Marks]

(A) 3.7×10^{-4} M

(B) 1.11×10^{-4} M

(C) 3.7×10^{-3} M

(D) 1.11×10^{-3} M`

Explanation:

$[H_3O]^+$ for a solution having pH = 3 is given by

$[H_3O]^+ = 1 \times 10^{-3}$ moles/litre [$\therefore [H_3O]^+ = 10^{-pH}$]

Similarly for solution having pH = 4,

$[H_3O]^+ = 1 \times 10^{-4}$ moles/litre and for pH = 5

$[H_3O]^+ = 1 \times 10^{-5}$ moles/litre

Let the volume of each solution in the mixture be 1 L ,then total volume of mixture solution

(1+1+1) L = 3 L

Total $[H_3O]^+$ ion present in mixture solution

$= (10^{-3} + 10^{-4} + 10^{-5})$ moles

Then $[H_3O]^+$ ions concentration of mixture solution

$= ((10^{-3} + 10^{-4} + 10^{-5}) / 3)$ M = (0.00111 / 3) M

$= 0.00037 \text{ M} = 3.7 \times 10^{-4}$ M.

Question 166.

The value of equilibrium constant of the reaction

$HI(g) \rightleftharpoons (1/2)\, H_2(g) + (1/2)\, I_2(g)$ is 8.0.

The equilibrium constant of the reaction

$H_2(g) + I_2(g) \rightleftharpoons 2HI(g)$ will be:

[4 Marks]

(A) 1 / 16

(B) 16

(C) 1 / 8

(D) 1 / 64

Explanation:

Given : Equilibrium constant (K_1) for the reaction:

$$HI(g) \overset{K_1}{\rightleftharpoons} \frac{1}{2} H_2(g) + \frac{1}{2} I_2(g); \quad K_1 = 8; \quad \ldots (i)$$

To find equilibrium constant for the following reaction:

$$H_2(g) + I_2(g) \rightleftharpoons 2HI(g); \; K_2 = ? \qquad \ldots . (ii)$$

multiply (i) by 2, we get

$$2HI(g) \rightleftharpoons H_2(g) + I_2(g); \qquad K_1 = 8^2 = 64. \quad \ldots (iii)$$

Now reverse equation (iii), we get

$$H_2(g) + I_2(g) \rightleftharpoons 2HI(g); \; K = \frac{1}{64} \qquad \ldots . (iv)$$

Equation (iv) is the same as the required equation (ii), thus

K_2 for equation (ii) is $\frac{1}{64}$

Question 167.

Equimolar solutions of the following were prepared in water separately. Which one of the solutions will record the highest pH?

[4 Marks]

(A) $CaCl_2$

(B) $MgCl_2$

(C) $BaCl_2$

(D) $SrCl_2$

Explanation:

The highest pH will be recorded by the most basic solution. The basic nature of hydroxides of alkaline earth metals increase as we move from Mg to Ba thus the solution of $BaCl_2$ in water will be most basic and so it will have highest pH.

Question 168.

The sequence of ionic mobility in aqueous solution is:

[4 Marks]

(A) $Na^+ > K^+ > Rb^+ > Cs^+$

(B) $Rb^+ > K^+ > Cs^+ > Na^+$

(C) $K^+ > Na^+ > Rb^+ > Cs^+$

(D) $Cs^+ > Rb^+ > K^+ > Na^+$

Explanation:

Smaller the ion, more is its ionic mobility in aqueous solution. Ionic radii of the given alkali metals is in the order $Na^+ < K^+ < Rb^+ < Cs^+$ and thus expected ionic mobility will be in the same order $Cs^+ < Rb^+ < K^+ < Na^+$. However, due to high degree of solvation (or hydration) and because of lower size or high charge density, the hydrated ion size follows the same order $Cs^+ > Rb^+ > K^+ > Na^+$ and thus, conductivity order is $Cs^+ > Rb^+ > K^+ > Na^+$.

Question 169.

The alkali metals form salt-like hydrides by the direct synthesis at elevated temperature. The thermal stability of these hydrides decreases in which of the following orders?

[4 Marks]

(A) LiH > NaH > KH > RbH > CsH

(B) NaH > LiH > KH > RbH > CsH

(C) KH > NaH > LiH > CsH > RbH

(D) CsH > RbH > KH > NaH > LiH

Explanation:

The stability of alkali metal hydrides decreases from Li to Cs. It is due to the fact that M–H bonds become weaker with increase in size of alkali metals as we move down the group from Li to Cs. Thus the order of stability of hydrides is

LiH > NaH > KH > RbH > CsH

Question 170.

In the hydrocarbon

$CH_3 - CH = CH - CH_2 - C \equiv CH$
6 5 4 3 2 1

The state of hybridization of carbons 1, 3, and 5 are in the following sequence:

[4 Marks]

(A) sp, sp^3, sp^2

(B) sp^2, sp, sp^3

(C) sp, sp^2, sp^3

(D) sp^3, sp^2, sp

Explanation:

C–1 is sp hybridized (C≡C)

C–3 is sp^3 hybridized (C–C)

C–5 is sp^2 hybridized (C=C)

Correct sequence: sp, sp^3, sp^2

Question 171.

[4 Marks]

Base strength of :
(A) $H_3C-\bar{C}H_2$ (B) $H_2C=\bar{C}H$ and
(C) $H-C\equiv\bar{C}$
is in the order of :

(A) (C) > (B) > (A)

(B) (B) > (A) > (C)

(C) (A) > (C) > (B)

(D) (A) > (B) > (C)

Explanation:

The amount of s-character in various hybrid orbitals is as follows.

sp = 50%, sp^2 = 33%, sp^3 = 25%

Therefore, s-character of the C–H bond in acetylene (sp) is greater than that of the C–H bond in alkene (sp^2 hybridized), which in turn has greater s-character of the C–H bond in alkanes. Thus, owing to a high s-character of the C–H bond in alkynes, the electrons constituting this bond are more strongly held by the carbon nucleus with the result the hydrogen present on such a carbon atom can be easily removed as a proton.

The acidic nature of three types of C–H bonds follows the order

≡C–H > =C–H > –C–H

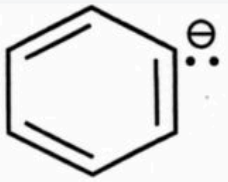
Further, as we know that conjugate base of a strong acid is a weak base, hence the correct order of basicity is:

$$H-C\equiv \overset{(-)}{C} < CH_2 = \overset{(-)}{C}H < \overset{(-)}{C}H_2-CH_3$$

Question 172.

The stability of carbanions in the following:

[4 Marks]

(a) $RC\equiv \overline{C}$ (b) 

(c) $R_2C=C\overline{H}$ (d) $R_3C-C\overline{H}_2$

is in the order of :

(A) (a) > (b) > (c) > (d)

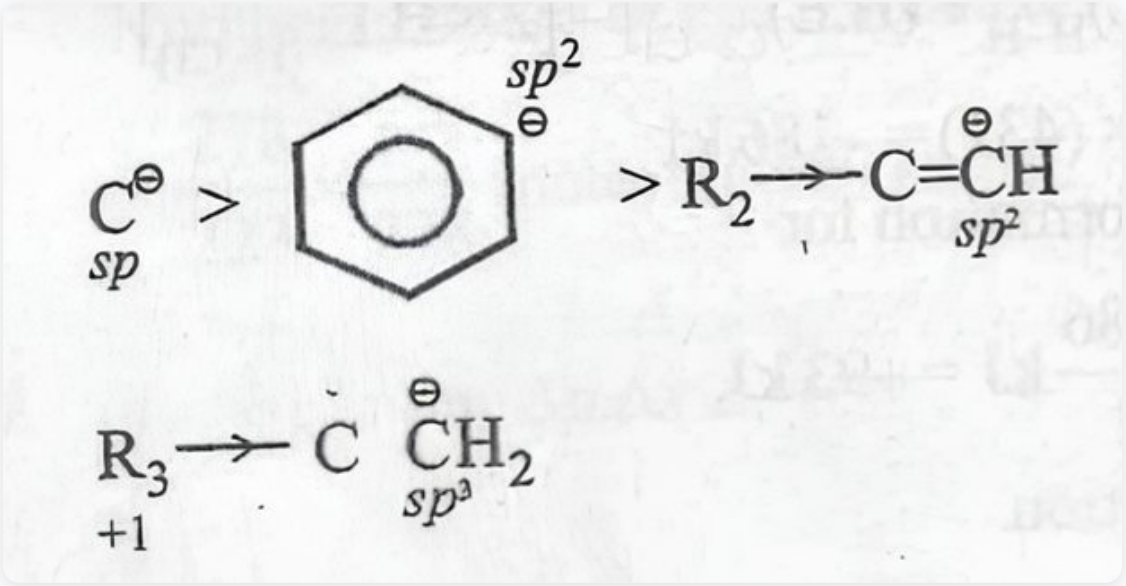
(B) (b) > (c) > (d) > (a)

(C) (d) > (b) > (c) > (a)

(D) (a) > (c) > (b) > (d)

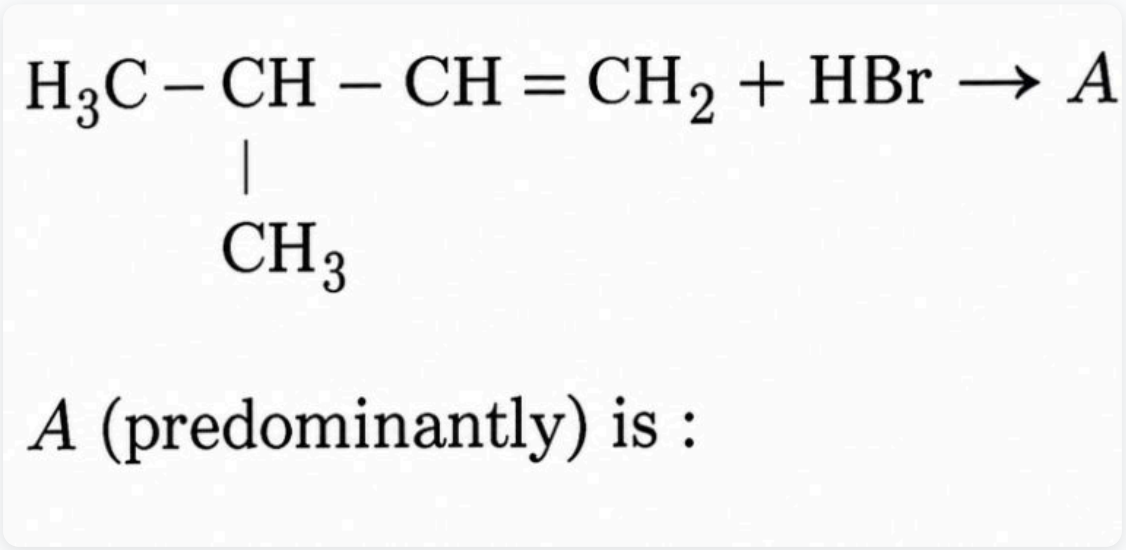
Explanation:

As 's' character increases, stability of carbanion increases. Electron donating groups decrease stability of carbanion. Hence, correct stability order:

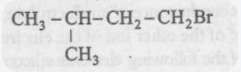


Question 173.

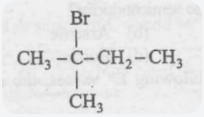
[4 Marks]



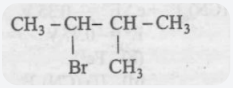
(A)



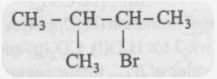
(B)



(C)

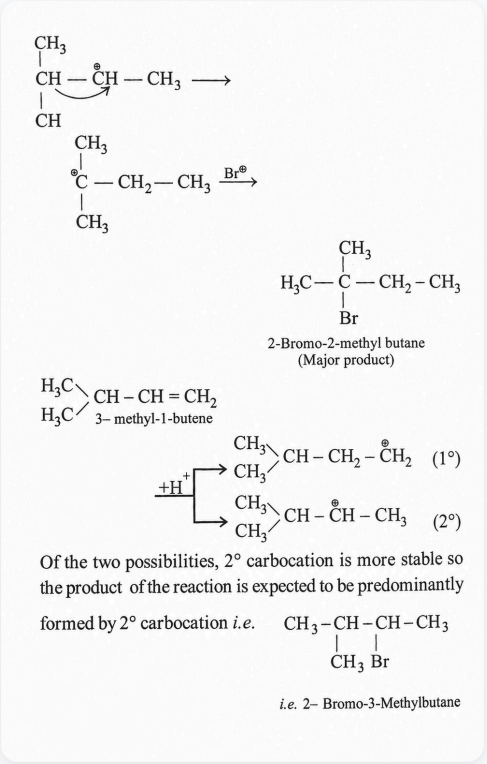


(D)



Explanation:

The formed 2° carbocation will rearrange to a more stable 3° carbocation as :

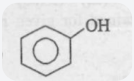


Question 174.

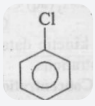
Which one of the following is most reactive towards electrophilic attack ?

[4 Marks]

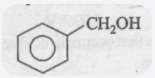
(A)



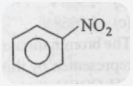
(B)



(C)

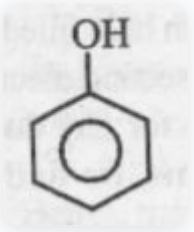


(D)



Explanation:

Out of the given compounds the most reactive towards nucleophilic attack is



Question 175.

Green chemistry means such reactions which

[4 Marks]

(A) are related to the depletion of ozone layer

(B) study the reactions in plants

(C) reduce the use and production of hazardous chemicals

(D) produce colour during reactions

Explanation:

Green chemistry are the reaction, which reduces the use of hazardous chemicals.

Question 176.

Percentage of free space in a body centred cubic unit cell is:

[4 Marks]

(A) 34%

(B) 28%

(C) 30%

(D) 32%

Explanation:

Percentage of occupied space in a bcc is 68%.

∴ Percentage of free space in a bcc is

(100 – 68) = 32%

Question 177.

If ‘a’ stands for the edge length of the cubic systems:

simple cubic, body centred cubic and face centred cubic, then the ratio of radii of the spheres in these systems will be respectively,

[4 Marks]

(A) a : √3a :√2a

(B) (1 / 2)a : √3a : (1 / √2)a

(C) (1 / 2)a : (√3 / 2)a : (√3 / 2)a

(D) (1 / 2)a : (√3 / 4)a : (1 / 2√2)a

Explanation:

Following generalization can be easily derived for various types of lattice arrangements in cubic cells between the edge length (a) of the cell and r the radius of the sphere.

For simple cubic : a = 2r or r = a / 2

For body centred cubic:

a = (4 / √3)r or r = (√3 / 4)a

For face centred cubic:

a = 2√2r or r = (1 / 2√2)a

Thus the ratio of radii of spheres will be simple: bcc : fcc

= a / 2 :(√3 / 4)a :(1 / 2√2)a

Question 178.

Kohlrausch’s law states that at:

[4 Marks]

(A) infinite dilution, each ion makes definite contribution to conductance of an electrolyte whatever be the nature of the other ion of the electrolyte.

(B) infinite dilution, each ion makes definite contribution to equivalent conductance of an electrolyte, whatever be the nature of the other ion of the electrolyte.

(C) finite dilution, each ion makes definite contribution to equivalent conductance of an electrolyte, whatever be the nature of the other ion of the electrolyte.

(D) infinite dilution each ion makes definite contribution to equivalent conductance of an electrolyte depending on the nature of the other ion of the electrolyte.

Explanation:

Kohlrausch's Law states that at infinite dilution, each ion migrates independently of its co-ion and contributes to the total equivalent conductance of an electrolyte a definite share which depends only on its own nature.

From this definition we can see that option (infinite dilution, each ion makes definite contribution to equivalent conductance of an electrolyte, whatever be the nature of the other ion of the electrolyte.) is the correct answer.

Question 179.

With which one of the following elements silicon should be doped so as to give p-type semiconductor?

[4 Marks]

(A) Selenium

(B) Boron

(C) Arsenic

(D) Germanium

Explanation:

The semiconductors formed by the introduction of impurity atoms containing one elecron less than the parent atoms of insulators are termed as p-type semiconductors.

Therefore, silicon containing 14 electrons has to be doped with boron containing 13 electrons to give a p-type semi-conductor.

Question 180.

On the basis of the following E° values, the strongest oxidizing agent is:

$[Fe(CN)_6]^{4-} \rightarrow [Fe(CN)_6]^{3-} + e^-$; $E° = -0.35$ V

$Fe^{2+} \rightarrow Fe^{3+} + e^-$ $E° = -0.77$ V

[4 Marks]

(A) Fe^{2+}

(B) $[Fe(CN)_6]^{4-}$

(C) Fe^{3+}

(D) $[Fe(CN)_6]^{3-}$

Explanation:

From the given data we find Fe^{3+}

($E°_{(Fe^{3+}/ Fe^{2+})}$ = + 0.77V) is strongest oxidizing agent.

Question 181.

Standard free energies of formation (in kJ/mol) at 298 K are –237.2, –394.4 and –8.2 for $H_2O(l)$, $CO_2(g)$ and pentane (g), respectively. The value of $E°_{cell}$ for the pentane–oxygen fuel cell is:

[4 Marks]

(A) 0.0968 V

(B) 1.0968 V

(C) 1.968 V

(D) 2.0968 V

Explanation:

Writing the equation for pentane–oxygen fuel cell at respective electrodes and over all reaction, we get

At Anode:

$C_5H_{12} + 10H_2O \rightarrow 5CO_2 + 32H^+ + 32e^-$

(pentane)

At Cathode:

$8O_2 + 32H^+ + 32e^- \rightarrow 16H_2O$

———————————————

Overall: $C_5H_{12} + 8O_2 \rightarrow 5CO_2 + 6H_2O$

Calculation of ΔG° for the above reaction

ΔG° = [5×(−394.4) + 6×(−237.2)] − [−8.2]

= −1972.0 − 1423.2 + 8.2 = −3387.0 kJ

= −3387000 Joules

From the overall equation we find n = 32

Using the relation ΔG° = −nFE°_cell and

substituting values we get

−3387000 = −32 × 96500 × E°_cell (F = 96500 C)

or E°_cell = 3387000 / (32×96500) = 3387000 / 3088000

= (3387 / 3088) V = 1.0968 V

Question 182.

The bromination of acetone that occurs in acid solution is represented by this equation:

$CH_3COCH_3(aq) + Br_2(aq) \rightarrow CH_3COCH_2Br(aq) + H^+(aq) + Br^-(aq)$

These kinetic data were obtained for given reaction concentrations.

[4 Marks]

Initial Concentrations, M		
$[CH_3COCH_3]$	$[Br_2]$	$[H^+]$
0.30	0.05	0.05
0.30	0.10	0.05
0.30	0.10	0.10
0.40	0.05	0.20
Initial rate, disappearance of Br_2, Ms^{-1}		
5.7×10^{-5}	5.7×10^{-5}	
1.2×10^{-4}	3.1×10^{-4}	

Base on these data, the rate equations is:

(A) Rate = $k[CH_3COCH_3][Br_2][H^+]$

(B) Rate = $k[CH_3COCH_3][Br_2][H^+]^2$

(C) Rate = $k[CH =COCH_3][Br_2]$

(D) Rate = $k[CH_3COCH_3][H^+]$

Explanation:

This reaction is autocatalyzed and involves complex calculation for concentration terms.

We can look at the above results in a simple way to find the dependence of reaction rate (i.e. rate of disappearance of Br_2).

From data (1) and (2) in which concentration of CH_3COCH_3 and H^+ remain unchanged and only the concentration of Br_2 is doubled, there is no change in rate of reaction. It means the rate of reaction is independent of concentration of Br_2.

Again from (2) and (3) in which (CH_3COCH_3) and (Br_2) remain constant but H^+ increases from 0.05 M to 0.10 i.e. doubled, the rate of reaction changes from 5.7×10^{-5} to 1.2×10^{-4} (or 12×10^{-5}), thus it also becomes almost doubled. It shows that rate of reaction is directly proportional to $[H^+]$. From (3) and (4), the rate should have doubled due to increase in conc of $[H^+]$ from 0.10 M to 0.20 M but the rate has changed from 1.2×10^{-4} to 3.1×10^{-4}. This is due to change in concentration of CH_3COCH_3 from 0.30 M to 0.40 M. Thus, the rate is directly proportional to $[CH_3COCH_3]$.

Rate = $k[CH_3COCH_3]^1[Br_2]^0[H^+]^1$

= $k[CH_3COCH_3][H^+]$

Rewriting the given data for the reaction

$$CH_3COCH_3(aq) + Br_2(aq) \xrightarrow{H^+} CH_3COCH_2Br(aq) + H^+(aq) + Br^-(aq)$$

S. No.	Initial concent -ration of CH_3COCH_3 in M	Initial concentr -ation of Br_2 in M	Initial concentr -ation of H^+ in M	Rate of disappearance of Br_2 in MS^{-1} i.e. $-\frac{d}{dt}[Br_2]$ or $\frac{dx}{dt}$
1	0.30	0.05	0.05	5.7×10^{-5}
2	0.30	0.10	0.05	5.7×10^{-5}
3	0.30	0.10	0.10	1.2×10^{-4}
4	0.40	0.05	0.20	3.1×10^{-4}

Question 183.

The rate constants k_1 and k_2 for two different reactions are 10^{16}. e_(-2000/T) and 10^{15}.e_(-1000/T), respectively. The temperature at which $k_1 = k_2$ is:

[4 Marks]

(A) (2000 / 2.303) K

(B) (1000 / 2.303) K

(C) 1000 K

(D) 2000 K

Explanation:

Given: $k_1 = 10^{16}$ e_(−2000/T) and $k_2 = 10^{15}$ .e_(-1000/T)

When $k_1 = k_2$ at temperature T, we have

10^{16}. e_(-2000/T) = 10^{15}. e_(−1000/T)

or 10×10^{15} e_(−2000/T) = 10^{15} e_(−1000/T)

or 10 .e_(−2000/T) = e_(−1000/T)

or ln 10 - (2000/T) = − (1000/T)

or ln 10 = 2000/T - 1000/T

or 2.303 log 10 = 1000/T

or 2.303 × 1 × T = 1000 [∵ log 10 = 1]

or T = (1000 / 2.303) K

Question 184.

The correct order of decreasing second ionisation enthalpy of Ti (22), V (23), Cr (24) and Mn (25) is:

[4 Marks]

(A) V > Mn > Cr > Ti

(B) Mn > Cr > Ti > V

(C) Ti > V > Cr > Mn

(D) Cr > Mn > V > Ti

Explanation:

Ti ; Z (22) is $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^2$

V ; Z (23) is $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^3$

Cr ; Z (24) is $1s^2 2s^2 2p^6 3s^2 3p^6 3d^5 4s^1$

Mn ; Z (25) is $1s^2 2s^2 2p^6 3s^2 3d^5 4s^2$

The Second electron in all the cases (except Cr) is taken out from 4s orbital and for Cr it is taken from half-filled 3d orbital. The force required for removal of second electron will be more for Mn than (except for Cr) due to having more positive charge. Based on this, we find the correct order Mn > V > Ti.

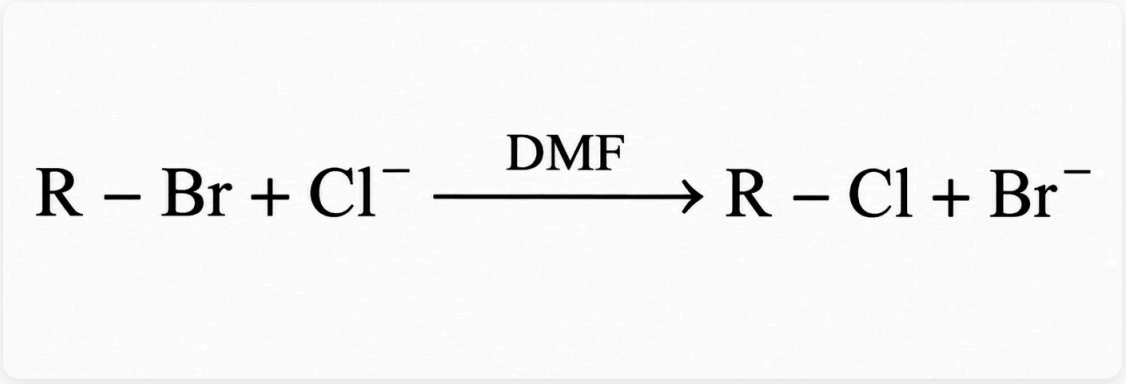
i.e. Cr > Mn > V > Ti.

Question 185.

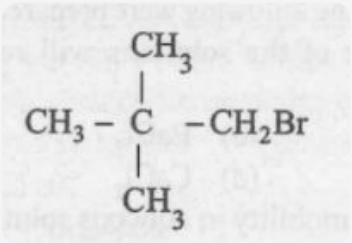
In an S_N2 substitution reaction of the type

which one of the following has the highest reactivity rate?

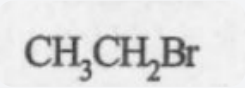
[4 Marks]



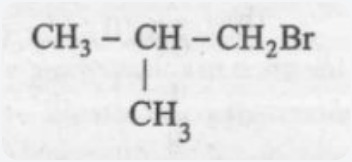
(A)



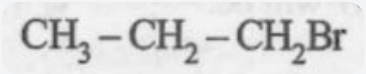
(B)



(C)

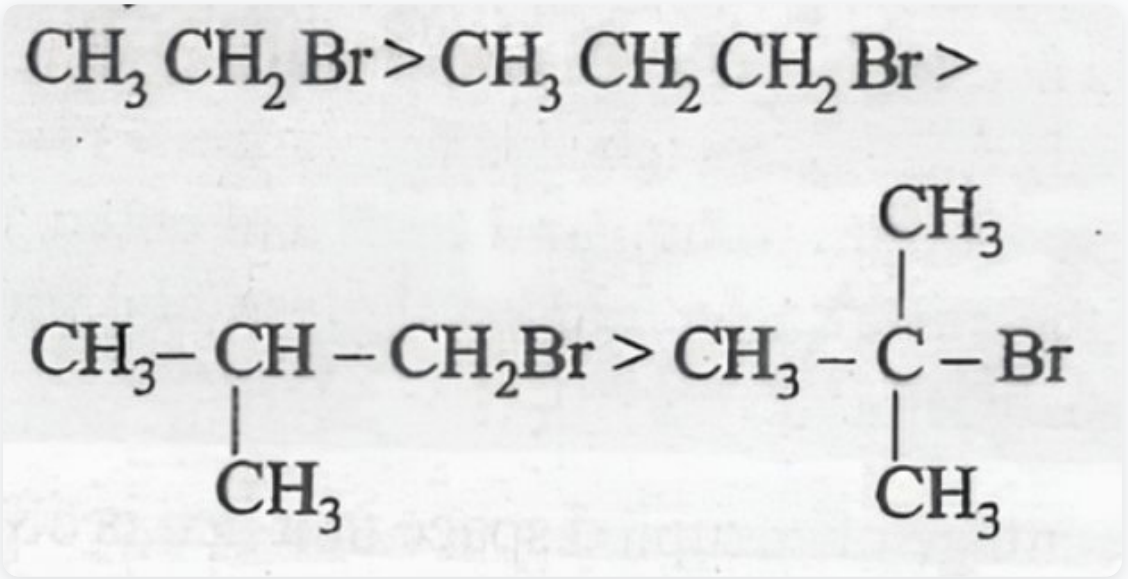


(D)



Explanation:

S_N2 mechanism is followed in case of primary and secondary halides i.e., S_N2 reaction is favoured by small groups on the carbon atom attached to halogens so reactivity order is:



Question 186.

How many stereoisomers does this molecule have?

$CH_3CH = CHCH_2CHBrCH_3$

[4 Marks]

(A) 4

(B) 2

(C) 6

(D) 8

Explanation:

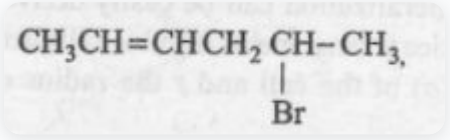
In the molecule

the number of stereoisomers is given by sum of geometrical isomers (because of presence of C = C) and optical isomers (because of presence of chiral carbon atom).

Number of geometrical isomers = 2 (one C = C is present).

Number of optical isomers = 2 (one chiral carbon atom).

Total number of stereoisomers = 2 + 2 = 4



Question 187.

A strong base can abstract an α-hydrogen from:

[4 Marks]

(A) alkene

(B) amine

(C) alkane

(D) ketone

Explanation:

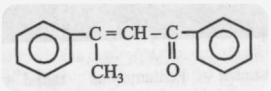
A strong base (OH^-) abstract an α — H atom from ketone which is acidic in nature.

Question 188.

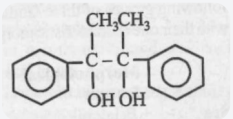
Acetophenone when reacted with a base, C_2H_5ONa, yields a stable compound which has the structure.

[4 Marks]

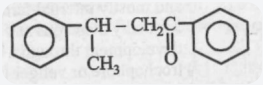
(A)



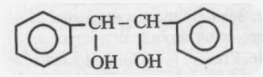
(B)



(C)

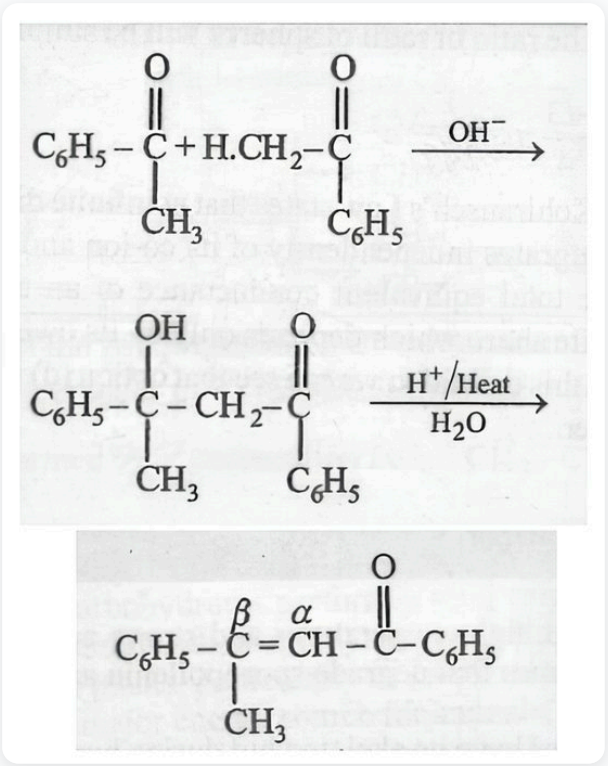


(D)



Explanation:

Thus, two molecules of acetophenone condense to form a β-hydroxy ketone which gets dehydrated in the presence of acid upon heating to form α, β-unsaturated compound.



Question 189.

The relative reactivities of acyl compounds towards nucleophilic substitution are in the order of:

[4 Marks]

(A) Acyl chloride > Acid anhydride > Ester > Amide

(B) Acid anhydride > Amide > Ester > Acyl chloride

(C) Ester > Acyl chloride > Amide > Acid anhydride

(D) Acyl chloride > Ester > Acid anhydride > Amide

Explanation:

The more the basic character of the leaving group, the lesser is the reactivity. The basic character follows the order:

$NH_2^- > OR^- > RCOO^- > Cl^-$

Hence, the relative reactivities of acyl compounds towards nucleophilic substitution follow the order Acyl halides > Acid anhydride > Ester > Amide.

Question 190.

Which one of the following is an amine hormone?

[4 Marks]

(A) Insulin

(B) Progesterone

(C) Oxypurin

(D) Thyroxine

Explanation:

Thyroxine is an amino hormone.

Question 191.

Which one of the following statement is not true?

[4 Marks]

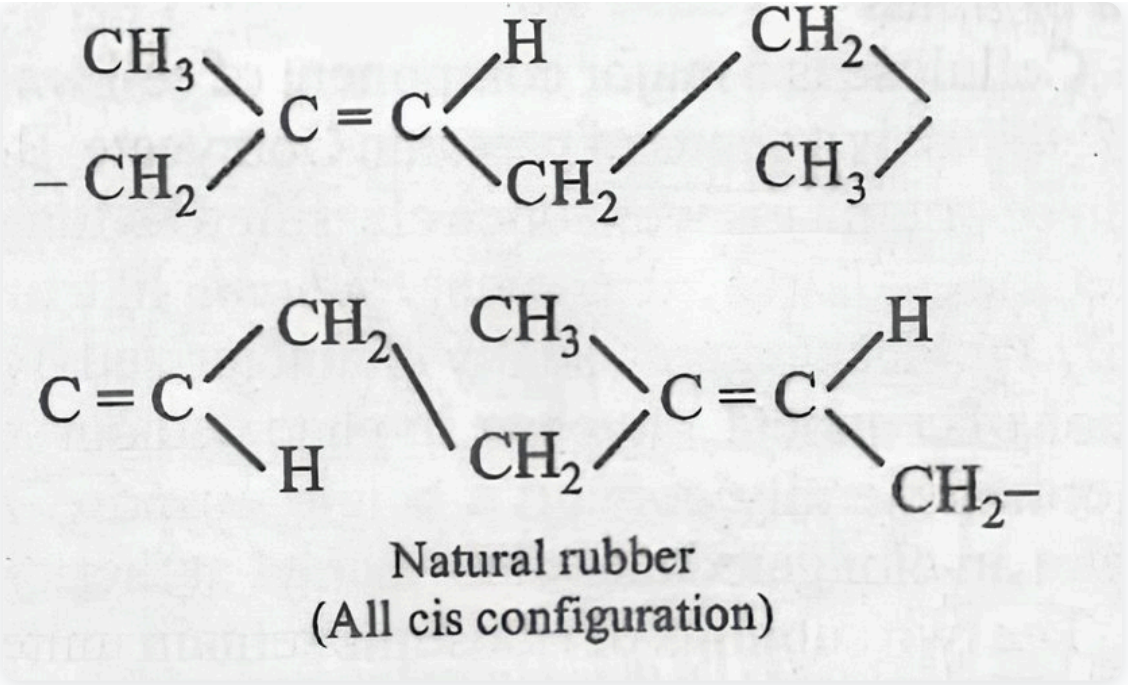
(A) Buna-S is a copolymer of butadiene and styrene

(B) Natural rubber is a 1,4-polymer of isoprene

(C) Natural rubber has the trans-configuration at every double bond

(D) In vulcanization the formation of sulphur bridges between different chains make rubber harder and stronger.

Explanation:



Question 192.

Which of the following complexes exhibits the highest paramagnetic behaviour?

where gly = glycine, en = ethylenediamine and bpy = bipyridyl moieties

(At.nos Ti = 22, V = 23, Fe = 26, Co = 27)

[4 Marks]

(A) $[V(gly)_2(OH)_2(NH_3)_2]^+$

(B) $[Fe(en)(bpy)(NH_3)_2]^{2+}$

(C) $[Ti(NH_3)_6]^{3+}$

(D) $[Co(ox)_2(OH)_2]^-$

Explanation:

More the number of unpaired electrons present in a complex, more is its paramagnetic behaviour.

(a) $\overset{x}{[V}\overset{-1}{(gly)_2}\overset{-1}{(OH)_2}\overset{0}{(NH_3)_2]^+}$

$x+2(-1)+2(-1)+2(0)=+1$

$\Rightarrow x=+5$

$V^{5+}=[Ar]$; no unpaired electron

(b) $\overset{x}{[Fe}\overset{0}{(en)}\overset{0}{(bpy)}\overset{0}{(NH_3)_2]^{2+}}$

$x=+2$

$Fe^{2+}=[Ar]\,3d^6$

But due to presence of strong field ligands, *en*, *bpy* and NH_3; the electrons will try to pair up. Thus, the complex will contain no unpaired electron.

(c) $\overset{e}{[Co}\overset{-2}{(Ox)_2}\overset{-1}{(OH)_2]^-}$

$x+2(-2)+2(-1)=-1 \Rightarrow x=+5$

$Co^{5+}=[Ar]\,3d^4$

Thus, 4 unpaired electrons.

(d) $\overset{x}{[Ti}\overset{0}{(NH_3)_6]^{3+}}$

$x=+3$

$Ti^{3+}=[Ar]\,3d^1$

Thus, 1 unpaired electron.

Question 193.

In which of the following coordination entities the magnitude Δ_0 (CFSE in octahedral field) will be maximum?

(At. No. Co = 27)

[4 Marks]

(A) $[Co(CN)_6]^{3-}$

(B) $[Co(C_2O_4)_3]^{3-}$

(C) $[Co(NH_3)_6]^{3+}$

(D) $[Co(H_2O)_6]^{3+}$

Explanation:

Out of the given ligands water, ammonia, cyanide and oxalate, maximum splitting will occur in case of cyanide (CN^-), i.e. the magnitude of Δ_0 will be maximum in case of $[Co(CN)_6]^{3-}$.

Question 194.

Which of the following statements is not correct?

[4 Marks]

(A) The number of carbon atoms in an unit cell of diamond is 4

(B) Molecular solids are generally volatile

(C) The fraction of the total volume occupied by the atoms in a primitive cell is 0.48

(D) The number of Bravais lattices in which a crystal can be categorized is 14

Explanation:

Diamond has ZnS type structure. So, number of atoms in a unit cell of diamond is 8.

Question 195.

Four diatomic species are listed below the different sequences. Which of these presents the correct order of their increasing bond order?

[4 Marks]

(A) NO < C_2^{2-} < O_2^- < $He2^+$

(B) C_2^{2-} < He_2^+ < NO < O_2^-

(C) O_2^- < NO < O_2^{2-} < He_2^+

(D) He_2^+ < O_2^- < NO < C_2^{2-}

Explanation:

He_2^+ < O_2^- < NO < C_2^{2-}

Question 196.

In DNA, the complimentary bases are

[4 Marks]

(A) Adenine and thymine, guanine and cytosine

(B) Adenine and guanine, thymine and cytosine

(C) Adenine and thymine, guanine and uracil

(D) Uracil and adenine, cytosine and guanine

Explanation:

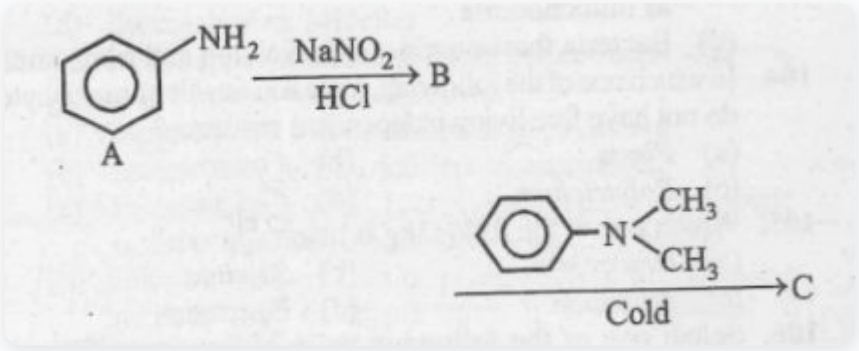
A = = = T

G ≡ ≡ ≡ C

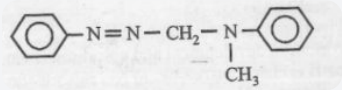
Question 197.

In a reaction of aniline a coloured product C was obtained The structure of C would be

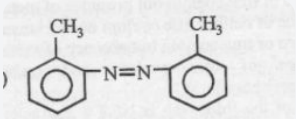
[4 Marks]



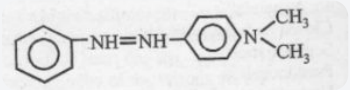
(A)



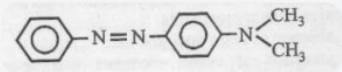
(B)



(C)

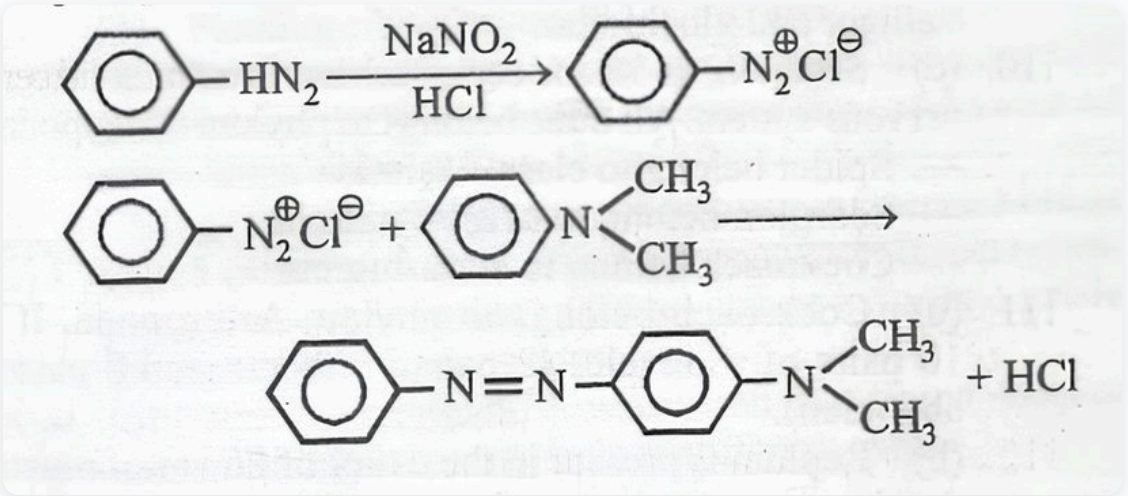


(D)



Explanation:

Diazotisation followed by electrophilic substitution at para position.



Question 198.

Which one of the following proved effective for biological control of nematodal disease in plants?

[4 Marks]

(A) Pisolithus tinctorius

(B) Pseudomonas cepacia

(C) Gliocladium virens

(D) Paecilomyces lilacinus

Explanation:

Paecilomyces is a cosmopolitan filamentous fungus which inhabits the soil, decaying plants, and food prod-ucts. Some species of Paecilomyces are isolated from insect. The telomorphs of Paecilomyces are classified in the genera Cilromocleista, Talaromyces. Paecilomyces is usually considered as a contaminant but may also cause infections in humans and animals.

Question 199.

Trichoderma harzianum has proved a useful microorgan-ism for

[4 Marks]

(A) reclamation of wastelands

(B) bioremediation of contaminated soils

(C) gene transfer in higher plants

(D) biological control of soil-borne plant pathogens

Explanation:

Trichoderma harzianum has proved a useful microor- 18 ganism for biological control of soil borne plant pathogens. Trichoderma harzianum is a fungus that is also used as a fun-gicide. It is used for foliar application, seed treatment and soil treatment for suppression of various disease causing fungal pathogens. Commercial biotechnological products such as 3 Tac have been useful for treatment of Botrytis, Fusarium, Penicillium sp. It is also used for manufacturing enzymes.

Question 200.

Modern detergents contain enzyme preparations of:

[4 Marks]

(A) Alkaliphiles

(B) Acidophiles

(C) Thermoacidophiles

(D) Thermophiles

Explanation:

Modern detergent contain enzymes preparations of alkaliphiles. Alkaliphiles live in alkaline environments like soda lakes, or alkaline soil. The pH level in alkaline substances is from about 9 to 11 on the pH scale. The al-kaliphiles keep an alkaline level of about pH 9 inside the cells, but the environment has a higher level of alkalinity. Alkaliphiles can keep the right pH level by pumping in hydrogen ions in out at the right rate.