

Chemical Reactions and Equations

Previous Year Questions

- Q.1.** Why does the color of copper sulphate solution change when an iron nail is dipped in it?
- Q.2.** A shiny brown colored element 'X' on heating in air becomes black in color. Name the element 'X' and the black colored compound formed. Write the chemical equation for the reaction.
- Q.3.** Why is respiration considered an exothermic reaction? Explain.
- Q.4.** What happens when: (Write balanced chemical equations)
- (a) Lead nitrate is heated in a boiling tube.
 - (b) Iron nails are dipped in copper sulphate solution.
 - (c) Barium chloride reacts with sodium sulphate solution.
- Q.5.** Define the following terms with one example of each:
- (i) Combination reaction
 - (ii) Decomposition reaction
 - (iii) Displacement reaction
- Q.6.** List two observations when zinc granules are added to dilute H_2SO_4 in a test tube. Write the chemical equation for the reaction.
- Q.7.** What is observed when a solution of sodium sulphate is added to a solution of barium chloride in a test tube? Write the chemical equation and name the type of reaction.
- Q.8.** List two safety precautions that should be followed while burning a magnesium ribbon in air. Mention two observations of this activity.
- Q.9.** (a) State the change in colour observed when silver chloride is exposed to sunlight. Name the type of chemical reaction that takes place.
(b) Write the chemical equation for the above reaction.

- Q.10.** (a) A student takes about 4 g of ferrous sulphate crystals in a dry boiling tube and heats the tube over the flame of a burner. State two observations.
- (b) Write the chemical equation for the reaction and name the type of reaction.